

INTRO TO IONIC BONDING

GUIDED NOTES



Ionic Bonding

All ions have a charge (lost electrons = positive charge, gained electrons = negative charge)

Write down the names of the following ions. Describe how they will attract ions.

POSITIVE IONS
Write down the name of the ion.

NEGATIVE IONS
Write down the name of the ion.

IONIC BONDING
Write down the definition of ionic bonding. Explain how it is formed.

Draw a dot-and-cross diagram for the following ions:

- Na^+
- Mg^{2+}
- O^{2-}
- Cl^-

Write down the name of the compound formed by the following ions:

- Na^+ and Cl^-
- Mg^{2+} and O^{2-}
- Al^{3+} and S^{2-}

Draw a dot-and-cross diagram for the following compounds:

- NaCl
- MgO
- Al_2S_3

IONIC BONDING

Name: _____

Directions: Answer the following.

Write W for metal and N for nonmetal for each of the following.

- _____ gains electrons when bonding
- _____ found on the left side of the periodic table
- _____ has just a few valence electrons
- _____ aluminum
- _____ has more than four valence electrons
- _____ loses electrons when bonding
- _____ found on the right side of the periodic table
- _____ bromine

Draw a Lewis structure for each of the following ionic compounds.

$\text{Strontium} + \text{Oxygen}$	$\text{Hydrogenium} + \text{Chlorine}$
$\text{Lithium} + \text{Fluorine}$	$\text{Calcium} + \text{Phosphorus}$

Self Assessment Qs

Classwork Worksheet



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