

## Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

### Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "em-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

### Model 1 – Ion Charges for Selected Elements

1	H <sup>+</sup>										
2	Li <sup>+</sup>	Be <sup>2+</sup>						N <sup>3-</sup>	O <sup>2-</sup>	F <sup>-</sup>	
3	Na <sup>+</sup>	Mg <sup>2+</sup>	Transition elements				Al <sup>3+</sup>		P <sup>3-</sup>	S <sup>2-</sup>	Cl <sup>-</sup>
4	K <sup>+</sup>	Ca <sup>2+</sup>	Fe <sup>2+</sup> Fe <sup>3+</sup>	Ni <sup>2+</sup> Ni <sup>3+</sup>	Cu <sup>+</sup> Cu <sup>2+</sup>	Zn <sup>2+</sup>					Br <sup>-</sup>
5	Rb <sup>+</sup>	Sr <sup>2+</sup>			Ag <sup>+</sup>			Sn <sup>2+</sup> Sn <sup>4+</sup>			I <sup>-</sup>
6		Ba <sup>2+</sup>				Hg <sub>2</sub> <sup>2+</sup> Hg <sup>2+</sup>		Pb <sup>2+</sup> Pb <sup>4+</sup>			
	← Cations →						← Anions →				

1. Based on the information in Model 1:
  - a. Identify three elements that form only one cation.  
*Li, Na, K, Rb, Be, Mg, Ca, Sr, Ba, Ag, Zn, Al (H according to this sheet, but technically not)*
  - b. Identify three elements that form only one anion.  
*N, P, O, S, F, Cl, Br, I*
  - c. Identify three elements that form more than one cation.  
*Fe, Ni, Cu, Hg, Sn, Pb*
  - d. In what region of the periodic table are these "multiple ion" elements usually located?  
*Center, transition elements (metals)*
2. Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.  
*K<sub>2</sub>S*
3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



# Ionic Bonds Packet Answer Key

**Sabine Zange**



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## **Table of Contents Ionic Bonds Packet Answer Key**

1. Understanding the eBook Ionic Bonds Packet Answer Key
  - The Rise of Digital Reading Ionic Bonds Packet Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds Packet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic Bonds Packet Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds Packet Answer Key
  - Personalized Recommendations
  - Ionic Bonds Packet Answer Key User Reviews and Ratings
  - Ionic Bonds Packet Answer Key and Bestseller Lists
5. Accessing Ionic Bonds Packet Answer Key Free and Paid eBooks
  - Ionic Bonds Packet Answer Key Public Domain eBooks

- Ionic Bonds Packet Answer Key eBook Subscription Services
- Ionic Bonds Packet Answer Key Budget-Friendly Options
- 6. Navigating Ionic Bonds Packet Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Ionic Bonds Packet Answer Key Compatibility with Devices
  - Ionic Bonds Packet Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ionic Bonds Packet Answer Key
  - Highlighting and Note-Taking Ionic Bonds Packet Answer Key
  - Interactive Elements Ionic Bonds Packet Answer Key
- 8. Staying Engaged with Ionic Bonds Packet Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ionic Bonds Packet Answer Key
- 9. Balancing eBooks and Physical Books Ionic Bonds Packet Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ionic Bonds Packet Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Bonds Packet Answer Key
  - Setting Reading Goals Ionic Bonds Packet Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Bonds Packet Answer Key
  - Fact-Checking eBook Content of Ionic Bonds Packet Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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