

9

Thermodynamics

1. Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
- (a) ΔS is negative and therefore, ΔH should be highly positive
(b) ΔS is negative and therefore, ΔH should be highly negative
(c) ΔS is positive and therefore, ΔH should be negative
(d) ΔS is positive and therefore, ΔH should also be highly positive
2. For the reaction, $X_2O_4(l) \longrightarrow 2XO_2(g)$, $\Delta U = 2.1 \text{ kcal}$, $\Delta S = 20 \text{ cal K}^{-1}$ at 300 K. Hence, ΔG is [2014]
- (a) 2.7 kcal (b) -2.7 kcal
(c) 9.3 kcal (d) -9.3 kcal
3. For a given exothermic reaction, K_p and K_p' are the equilibrium constants at temperatures T_1 and T_2 , respectively. Assuming that heat of reaction is constant in temperature range between T_1 and T_2 , it is readily observed that [2014]
- (a) $K_p > K_p'$
(b) $K_p < K_p'$
(c) $K_p = K_p'$
(d) $K_p = \frac{1}{K_p'}$
4. A reaction having equal energies of activation for forward and reverse reactions has [2013]
- (a) $\Delta S = 0$ (b) $\Delta G = 0$
(c) $\Delta H = 0$ (d) $\Delta H = \Delta G = \Delta S = 0$
5. In which of the following reactions, standard reaction entropy changes (ΔS°) is positive and standard Gibbs energy change (ΔG°) decreases sharply with increasing temperature? [2012]
- (a) $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$
(b) $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$
(c) $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$
(d) $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
6. The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
- (a) 10.52 cal/mol K (b) 21.04 cal/mol K
(c) 5.260 cal/mol K (d) 0.526 cal/mol K
7. Standard enthalpy of vaporisation $\Delta_{\text{vap}}H^\circ$ for water at 100°C is 40.66 kJ mol⁻¹. The internal energy of vaporisation of water at 100°C (in kJ mol⁻¹) is (assume water vapour to behave like an ideal gas). [2012]
- (a) + 37.56 (b) - 43.76
(c) + 43.76 (d) + 40.66
8. If the enthalpy change for the transition of liquid water to steam is 30 kJ mol⁻¹ at 27°C, the entropy change for the process would be [2011]
- (a) 1.0 J mol⁻¹K⁻¹
(b) 0.1 J mol⁻¹K⁻¹
(c) 100 J mol⁻¹K⁻¹
(d) 10 J mol⁻¹K⁻¹

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