

# IDEAL GAS LAW

Name: \_\_\_\_\_

Use the Ideal Gas Law below to solve the following problems.

$PV = nRT$  where  $P$  is pressure in atmospheres,  
 $V$  is volume in liters,  
 $n$  is number of moles of gas,  
 $R$  is Ideal Gas Constant  
 $0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$ ,  
 $T$  is Kelvin temperature.

moles of gas =  $n$   
 volume =  $V$   
 pressure =  $P$

1. What volume (in liters) will occupy 0.500 moles of  $\text{CO}_2$  at 1.00 atm and  $27^\circ\text{C}$ ?  
11.2 L  $(PV = nRT \rightarrow V = nRT/P)$

2. What volume will 0.250 moles of nitrogen occupy at 1.00 atm and  $27^\circ\text{C}$ ?  
5.62 L  $(PV = nRT \rightarrow V = nRT/P)$

3. What pressure will be exerted by 25.0 g of  $\text{CO}_2$  at a temperature of  $27^\circ\text{C}$  in a volume of 0.500 L?  
11.8 atm  $(PV = nRT \rightarrow P = nRT/V)$

4. At what temperature will 0.500 g of  $\text{O}_2$  exert a pressure of 1000. torr in a volume of 0.100 L?  
11.2 K  $(PV = nRT \rightarrow T = PV/nR)$

5. What is the density of air,  $\text{molar mass} = 29 \text{ g/mol}$  at 1.00 atm and  $27^\circ\text{C}$ ?  
1.18 g/L  $(PV = nRT \rightarrow n = PV/RT \rightarrow \text{mass} = n \cdot \text{molar mass} \rightarrow \text{density} = \text{mass}/V)$

6. If the density of nitrogen is 1.15 g/L at 1.00 atm and  $27^\circ\text{C}$ , what is its molar mass?  
28.0 g/mol  $(PV = nRT \rightarrow n = PV/RT \rightarrow \text{mass} = n \cdot \text{molar mass} \rightarrow \text{density} = \text{mass}/V)$

7. How many moles of nitrogen will occupy a volume of 0.500 L at a total pressure of 1.00 atm?  
0.020 moles  $(PV = nRT \rightarrow n = PV/RT)$

8. What volume will 0.250 moles of  $\text{O}_2$  at a temperature of  $27^\circ\text{C}$  after 0.500 atm?  
5.62 L  $(PV = nRT \rightarrow V = nRT/P)$

9. Find the number of grams of  $\text{CO}_2$  that will be produced if 250.0 mL of a mixture of 25%  $\text{CO}_2$  at a temperature of  $27^\circ\text{C}$ .  
0.250 g  $(PV = nRT \rightarrow n = PV/RT \rightarrow \text{mass} = n \cdot \text{molar mass})$

10. A 100.0 mL gas flask is filled with 0.500 g of  $\text{O}_2$  at a pressure of 1.00 atm. What is the temperature of the gas?  
11.2 K  $(PV = nRT \rightarrow T = PV/nR)$

moles of gas =  $n$

# Gas Law Packet Answer Key

**Rachel S Tattersall**



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**Chemistry Homework** Frank Schaffer Publications, Joan DiStasio, 1996-03 Includes the periodic table writing formulas balancing equations stoichiometry problems and more

**Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

**Building Block** Prentice-Hall Staff, 1994

*Practical English* Vicki L. Hackett, Paul C. Dalmas, 1987 This book includes daily lesson plans and supplemental materials for a course in vocational

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