

# Gas Laws Problem Packet

**Gas Laws Packet**

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

**Essentials**

Standard Constant:

Unit	Standard Pressure
atm	1.013
mmHg	760
cmHg	76
torr	760
atm	1.013

**22.4 L = 1 mol at STP**

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law:	$P_1V_1 = P_2V_2$
Charles's Law:	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law:	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Avogadro's Law:	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law:	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Graham's Law:	$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$

The constant R (used in the other gas laws) depends on the unit of Pressure:

- 1 atm,  $R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$
- 1 atm,  $R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$

**Boyle's Law**

- Temperature must always be in Kelvin ( $K = ^\circ C + 273$ )
- If you are given two different units of pressure, CONVERT (and)
- Create two columns before solving, initial and final, write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (for Ideal Gas Law, write what you're given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE (MOL) (shown)
- Graham's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

**Charles' Law  $V \propto T$**

Directions: Complete each of the problems below. Show your work for all of the problems.

- The temperature inside my refrigerator is about 4°C. If I place it below the freezer, how will the temperature of the air inside the refrigerator change? Will it expand or contract? Why?
- A balloon is inflated to a volume of 1.0 L at 27°C. What will its volume be at 0°C?

**Gay-Lussac's Law  $P \propto T$**

- A gas is held at a constant volume. The pressure is 1.0 atm at 27°C. What will the pressure be at 100°C?
- A gas is held at a constant volume. The pressure is 1.0 atm at 27°C. What will the pressure be at 0°C?

**Avogadro's Law  $V \propto n$**

- A gas is held at a constant temperature and pressure. The volume is 1.0 L and the number of moles is 1.0 mol. What will the volume be if the number of moles is 2.0 mol?
- A gas is held at a constant temperature and pressure. The volume is 1.0 L and the number of moles is 1.0 mol. What will the volume be if the number of moles is 0.5 mol?

**Combined Gas Law  $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$**

- A gas is held at a constant temperature and pressure. The volume is 1.0 L, the pressure is 1.0 atm, and the temperature is 27°C. What will the volume be if the pressure is 2.0 atm and the temperature is 100°C?
- A gas is held at a constant temperature and pressure. The volume is 1.0 L, the pressure is 1.0 atm, and the temperature is 27°C. What will the volume be if the pressure is 2.0 atm and the temperature is 0°C?

**Graham's Law  $\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$**

- Hydrogen gas effuses through a porous barrier 9.7 times faster than an unknown gas. What is the molecular mass of the unknown gas?
- Hydrogen gas effuses through a porous barrier 3.16 times faster than an unknown gas. What is the molecular mass of the unknown gas?

All Gas Laws Included!

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# Gas Law Practice Homework Packet Key

**F Rizvi**



**Gas Law Practice Homework Packet Key:**

## Reviewing **Gas Law Practice Homework Packet Key**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Gas Law Practice Homework Packet Key**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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