

100% Grade Chemistry - Class Lesson (Major Exponent) - 1997

<p>1. In a chemical reaction, the <b>total mass of the reactants is equal to the total mass of the products.</b></p> <p>Example: <math>2H_2 + O_2 \rightarrow 2H_2O</math></p> <p>Mass of <math>H_2</math> = 4g Mass of <math>O_2</math> = 32g Total mass of reactants = 36g</p> <p>Mass of <math>H_2O</math> = 36g</p>	<p>2. In a chemical reaction, the <b>total number of atoms of each element is conserved.</b></p> <p>Example: <math>2H_2 + O_2 \rightarrow 2H_2O</math></p> <p>Number of H atoms = 4 Number of O atoms = 2</p> <p>Number of H atoms in <math>H_2O</math> = 4 Number of O atoms in <math>H_2O</math> = 2</p>	<p>3. The <b>total mass of a substance is conserved in a closed system.</b></p> <p>Example: <math>2H_2 + O_2 \rightarrow 2H_2O</math></p> <p>Mass of <math>H_2</math> = 4g Mass of <math>O_2</math> = 32g Total mass = 36g</p>
<p>4. The <b>total mass of a substance is conserved in a closed system.</b></p> <p>Example: <math>2H_2 + O_2 \rightarrow 2H_2O</math></p> <p>Mass of <math>H_2</math> = 4g Mass of <math>O_2</math> = 32g Total mass = 36g</p>	<p>5. The <b>total mass of a substance is conserved in a closed system.</b></p> <p>Example: <math>2H_2 + O_2 \rightarrow 2H_2O</math></p> <p>Mass of <math>H_2</math> = 4g Mass of <math>O_2</math> = 32g Total mass = 36g</p>	<p>6. The <b>total mass of a substance is conserved in a closed system.</b></p> <p>Example: <math>2H_2 + O_2 \rightarrow 2H_2O</math></p> <p>Mass of <math>H_2</math> = 4g Mass of <math>O_2</math> = 32g Total mass = 36g</p>
<p>7. The <b>total mass of a substance is conserved in a closed system.</b></p> <p>Example: <math>2H_2 + O_2 \rightarrow 2H_2O</math></p> <p>Mass of <math>H_2</math> = 4g Mass of <math>O_2</math> = 32g Total mass = 36g</p>	<p>8. The <b>total mass of a substance is conserved in a closed system.</b></p> <p>Example: <math>2H_2 + O_2 \rightarrow 2H_2O</math></p> <p>Mass of <math>H_2</math> = 4g Mass of <math>O_2</math> = 32g Total mass = 36g</p>	<p>9. The <b>total mass of a substance is conserved in a closed system.</b></p> <p>Example: <math>2H_2 + O_2 \rightarrow 2H_2O</math></p> <p>Mass of <math>H_2</math> = 4g Mass of <math>O_2</math> = 32g Total mass = 36g</p>

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