



GAS VOLUMES 1

$$\text{Volume (dm}^3\text{)} = 24 \times \text{moles}$$



The volume must be in dm^3 (there are 1000 cm^3 in 1 dm^3).

$$\text{vol in dm}^3 = \frac{\text{vol in cm}^3}{1000}$$

Give all answers to 3 significant figures.

1 Find the volume of the following gases (measured at room temperature and pressure).

- a 4.00 moles of oxygen (O_2) $\text{volume} = 4.00 \times 24 = 96.0 \text{ dm}^3$
- b 0.250 moles of methane (CH_4) $\text{volume} = 0.25 \times 24 = 6.00 \text{ dm}^3$
- c 15.0 g of argon (Ar) $\text{moles Ar} = \frac{15}{40} = 0.375 \text{ mol}$ $\text{volume} = 0.375 \times 24 = 9.00 \text{ dm}^3$
- d 0.220 g of carbon dioxide (CO_2) $\text{moles CO}_2 = \frac{0.220}{44} = 0.00500 \text{ mol}$ $\text{volume} = 0.00500 \times 24 = 0.120 \text{ dm}^3$

2 Find the number of moles of the following gases (measured at room temperature and pressure).

- a 48.0 dm^3 of carbon monoxide (CO) $\text{moles CO} = \frac{48.0}{24} = 2.00 \text{ mol}$
- b 1.20 dm^3 of hydrogen (H_2) $\text{moles H}_2 = \frac{1.20}{24} = 0.0500 \text{ mol}$
- c 360 cm^3 of oxygen (O_2) $\text{moles O}_2 = \frac{360}{24000} = 0.01500 \text{ mol}$

3 Find the mass of the following gases (measured at room temperature and pressure).

- a 7.20 dm^3 of ammonia (NH_3) $\text{moles NH}_3 = \frac{7.20}{24} = 0.300 \text{ mol}$ $\text{mass} = 17 \times 0.300 = 5.10 \text{ g}$
- b 480 cm^3 of nitrogen (N_2) $\text{moles N}_2 = \frac{480}{24000} = 0.0200 \text{ mol}$ $\text{mass} = 28 \times 0.0200 = 0.560 \text{ g}$
- c 100 cm^3 of oxygen (O_2) $\text{moles O}_2 = \frac{100}{24000} = 0.00417 \text{ mol}$ $\text{mass} = 32 \times 0.00417 = 0.133 \text{ g}$

Gas Variable Model 1 Answer Key

Michael Hoelscher



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