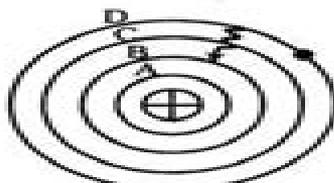


## Chapter-2: Qualitative Chemistry

Ques. ► 1



$$R_H = 109678 \text{ cm}^{-1}$$

[D.B., D.J.B., S.B., J.B.-18]

- What is the law of mass action?
- pH of pure water is 7 - Explain.
- Calculate the number of orbital in the last energy level of the above stem.
- If the electron of the stem is transferred into shell B, then what will be the color of the emitted ray? Explain mathematically.

### Answer to the question no. 1

At a fixed temperature, at a fixed time, the rate of any reaction is proportional to the active mass (i.e. molar concentration or partial pressure) of its reactant at that time.

The pH of any solution depends on the presence of molar concentration  $H^+$  and  $OH^-$  in that solution. As the concentration of  $[H^+]$  and  $[OH^-]$  produced due to the dissociation of pure water are almost same, the equation of ionic product is :

$$[H^+][OH^-] = 10^{-14}$$

$$\text{Or, } [H^+][H^+] = 10^{-14} \text{ [As, } [H^+] = [OH^-]]$$

$$\therefore [H^+] = 10^{-7}$$

Now, taking  $-\log$  on both sides,

$$-\log [H^+] = -\log 10^{-7}$$

$$\text{Or, pH} = 7$$

i.e., the pH of pure water is  $\text{pH} = 7$

So it can be said that, as the concentration of the ions produced by the dissociation of pure water are same, the pH of pure water is 7.

The principle quantum number in the fourth energy level is,  $n = 4$ .

If the principle quantum number is  $n$ , then the value of subsidiary quantum number is from zero (0) to  $(n-1)$ .

$$\therefore l = \text{from } 0 \text{ to } (n-1).$$

$$= \text{from } 0 \text{ to } (4-1).$$

$$= \text{from } 0 \text{ to } 3$$

$$= 0, 1, 2, 3.$$

Again, magnetic quantum number,  $m = -l$  to  $0$  to  $+l$ .

$$\therefore \text{When, } l = 0, m = 0, \text{ no. of orbital} = 1$$

$$l = 1, m = -1, 0, +1, \text{ no. of orbital} = 3$$

$$l = 2, m = -2, -1, 0, +1, +2, \text{ no. of orbital} = 5$$

$$l = 3, m = -3, -2, -1, 0, +1, +2, +3, \text{ no. of orbital} = 7$$

$$\therefore \text{When, } l = 0, 1, 2, 3, \text{ then the total number of orbital} = 16$$

We know that in a principle energy level, the total number of magnetic quantum number is the number of orbital in energy level. As in this case the total number of magnetic quantum number is 16, so the total number of orbital will be 16.

In the stem, the electron is positioned in principle energy level D i.e.  $n_2 = 4$  and after changing energy level it comes to the principle energy level B. i.e.  $n_1 = 2$ .

We know,

$$\frac{1}{\lambda} = R_H \left( \frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

$$\text{Or, } \frac{1}{\lambda} = 109678 \left( \frac{1}{2^2} - \frac{1}{4^2} \right)$$

$$\text{Or, } \frac{1}{\lambda} = 20564.625 \text{ cm}^{-1}$$

$$\text{Or, } \lambda = 4.862 \times 10^{-5} \text{ cm}$$

$$\text{Or, } \lambda = 4.862 \times 10^{-5} \times 10^7 \text{ nm [1 cm} = 10^7 \text{ nm]}$$

$$\therefore \lambda = 486.27 \text{ nm}$$

So, the determined wavelength is 486.27nm which is in the range of the color sky blue of wavelength 450 – 500 nm. So the color of the emitted ray of the mentioned electron will be sky blue.

Ques. ► 2



i. Solubility of  $MN_2 = 0.0002 \text{ mol L}^{-1}$

ii.  $K_{sp}$  of  $MY_2 = 1.85 \times 10^{-8}$

[D.B., D.J.B., S.B., J.B.-18]

- What is sigma bond?
- $Al_2O_3$  is an amphoteric oxide— Explain.
- Determine the solubility product of the salt in pot A.
- Will there be any ppt of  $MY_2$  in pot C? Explain mathematically.

### Answer to the question no. 2

The bond created due to face to face overlapping of the two valence orbitals during the creation of covalent bond is known as sigma bond.

The oxides which can act both as acid and base are known as amphoteric oxides. Here  $Al_2O_3$  reacts with both acid and base and produces salt. So, according to the characteristics  $Al_2O_3$  is an amphoteric oxide.



# Hsc Chemistry First Chapter Notes

**JS Bruner**



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