



<b>Activity B:</b> Classifying reactions	<u>Get the Gizmo ready:</u> <ul style="list-style-type: none"> <li>Turn off Show summary and Show histograms.</li> </ul>	
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**Introduction:** Chemical equations show how **compounds** and **elements** react with one another. An element is a substance consisting of one kind of atom, such as aluminum (Al) or oxygen gas ( $\text{O}_2$ ). A compound is a substance made of more than one kind of atom, such as water ( $\text{H}_2\text{O}$ ) or table salt (NaCl).

**Question:** How are chemical reactions classified?

1. **Match:** Most chemical reactions can be classified as one of four types. Using the chemical equations in the Gizmo as a guide, match the following definitions to the type of reaction.

B	One reactant is broken down into two or more products.
E	A fuel combines with oxygen to produce carbon dioxide ( $\text{CO}_2$ ) and water ( $\text{H}_2\text{O}$ ).
A	Two or more reactants combine to form one product.
D	Two compounds react to form two different compounds.
C	A compound reacts with an element to form a new compound and a different element.

A. <b>Synthesis</b>
B. <b>Decomposition</b>
C. <b>Single replacement</b>
D. <b>Double replacement</b>
E. <b>Combustion</b>

2. **Practice:** Balance each of the chemical equations below. (Some equations may already be in balance.) In the space to the right, classify the reaction as a *synthesis*, *decomposition*, *single replacement*, *double replacement*, or *combustion*.

A.	<input type="text" value="1"/>	$\text{AgNO}_3 +$	<input type="text" value="1"/>	$\text{KCl} \rightarrow$	<input type="text" value="1"/>	$\text{AgCl} +$	<input type="text" value="1"/>	$\text{KNO}_3$	<input type="text" value="Two replacements"/>
B.	<input type="text" value="1"/>	$\text{H}_2\text{O} +$	<input type="text" value="1"/>	$\text{SO}_3 \rightarrow$	<input type="text" value="1"/>	$\text{H}_2\text{SO}_4$			<input type="text" value="Synthesis"/>
C.	<input type="text" value="2"/>	$\text{KI} +$	<input type="text" value="1"/>	$\text{Cl}_2 \rightarrow$	<input type="text" value="2"/>	$\text{KCl} +$	<input type="text" value="1"/>	$\text{I}_2$	<input type="text" value="Single replacement"/>
D.	<input type="text" value="2"/>	$\text{NaHCO}_3 \rightarrow$	<input type="text" value="1"/>	$\text{Na}_2\text{CO}_3 +$	<input type="text" value="1"/>	$\text{H}_2\text{O} +$	<input type="text" value="1"/>	$\text{CO}_2$	<input type="text" value="Decomposition"/>
E.	<input type="text" value="1"/>	$\text{Zn} +$	<input type="text" value="2"/>	$\text{HCl} \rightarrow$	<input type="text" value="1"/>	$\text{ZnCl}_2 +$	<input type="text" value="1"/>	$\text{H}_2$	<input type="text" value="Single replacement"/>

# Gizmo Balancing Chemical Equations Answer Key

**Marcel A. Müller**



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