

Name _____ Sex _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data in the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , H_2 , and H_2O are 32.00 g/mol, 2.016 g/mol, and 18.02 g/mol, respectively.

Amount of O_2	Amount of H_2	Amount of H_2O	Limiting Reactant	Amount of Mass of Excess Reactant
1 molecule	1 molecule	1 molecule	none	none
2 molecules	2 molecules	2 molecules	H_2	1 molecule O_2
2 molecules	3 molecules	3 molecules	O_2	1 molecule H_2
1 mol	2 mol	2 mol	H_2	1 mol O_2
2 mol	3 mol	3 mol	O_2	1 mol H_2
1 mol	3 mol	2 mol	H_2	1 mol O_2
1 mol	4 mol	2 mol	H_2	2 mol O_2
1 mol	5 mol	2 mol	H_2	3 mol O_2
1 mol	6 mol	2 mol	H_2	4 mol O_2
1 mol	7 mol	2 mol	H_2	5 mol O_2
1 mol	8 mol	2 mol	H_2	6 mol O_2
1 mol	9 mol	2 mol	H_2	7 mol O_2
1 mol	10 mol	2 mol	H_2	8 mol O_2
1 mol	11 mol	2 mol	H_2	9 mol O_2
1 mol	12 mol	2 mol	H_2	10 mol O_2
1 mol	13 mol	2 mol	H_2	11 mol O_2
1 mol	14 mol	2 mol	H_2	12 mol O_2
1 mol	15 mol	2 mol	H_2	13 mol O_2
1 mol	16 mol	2 mol	H_2	14 mol O_2
1 mol	17 mol	2 mol	H_2	15 mol O_2
1 mol	18 mol	2 mol	H_2	16 mol O_2
1 mol	19 mol	2 mol	H_2	17 mol O_2
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1 mol	91 mol	2 mol	H_2	89 mol O_2
1 mol	92 mol	2 mol	H_2	90 mol O_2
1 mol	93 mol	2 mol	H_2	91 mol O_2
1 mol	94 mol	2 mol	H_2	92 mol O_2
1 mol	95 mol	2 mol	H_2	93 mol O_2
1 mol	96 mol	2 mol	H_2	94 mol O_2
1 mol	97 mol	2 mol	H_2	95 mol O_2
1 mol	98 mol	2 mol	H_2	96 mol O_2
1 mol	99 mol	2 mol	H_2	97 mol O_2
1 mol	100 mol	2 mol	H_2	98 mol O_2

balanced equation mole ratio = 2 mol H_2O /1 mol O_2

16.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.5121 mol O_2

20.00 g H_2 = 1 mol H_2 /2.016 g H_2 = 9.920 mol H_2

actual mole ratio = 9.920 mol H_2 /0.5121 mol O_2 = 19.37 mol H_2 /1 mol O_2

Because the actual mole ratio of H_2 / O_2 is larger than the balanced equation mole ratio of H_2 / O_2 , there is an excess of H_2 ; O_2 is the limiting reactant.

Mass of H_2 used = 0.5121 mol O_2 \times 2 mol H_2 /1 mol O_2 = 1.024 mol H_2

0.5121 mol O_2 \times 32.00 g O_2 /1 mol O_2 = 16.39 g O_2

Mass of H_2O produced = 0.5121 mol O_2 \times 2 mol H_2O /1 mol O_2 = 1.024 mol H_2O

Excess H_2 = 20.00 g H_2 - 1.024 g H_2 = 18.98 g H_2

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CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.42 g of the oxide from a 1.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_3N_2 are 69.72 g/mol and 177.00 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield volume = 7.42 g Ga_3N_2
2. mass of reactant = 1.00 g Ga
3. number of moles of reactant = 1.00 g Ga \times $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ = 7.00 g Ga = 1 mol Ga
4. number of moles of product = 0.00 mol Ga \times $\frac{2 \text{ mol Ga}_3\text{N}_2}{1 \text{ mol Ga}}$ = 0.00 mol Ga_3N_2
5. theoretical yield = 0.000 mol Ga_3N_2 \times $\frac{177.00 \text{ g Ga}_3\text{N}_2}{1 \text{ mol Ga}_3\text{N}_2}$ = correct
6. percent yield = $\frac{7.42 \text{ g Ga}_3\text{N}_2}{177.00 \text{ g Ga}_3\text{N}_2} \times 100$ = 4.19%

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