

Matter and Change, pp. 1-4

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|----|----|---|
| 2. | 5. | c |
| 3. | 5. | c |
| 4. | 7. | c |
| 9. | 3. | a |
10. mass and volume
 11. group or family
 12. homogeneous
 13. products
 14. metal
 15. solid state
 16. chemical properties
 17. atom
 18. compound
 19. nonmetal
 20. mixture
 21. pure substance
 22. basic
 23. metalloid
 24. reactants: carbon and oxygen; product: carbon dioxide
 25. reactant: mercury(II) oxide; products: mercury and oxygen
 26. f
 27. e
 28. c
 29. a
 30. b
 31. i
 32. d
 33. h
 34. e
 35. physical change
 36. chemical change
 37. physical change
 38. physical change
 39. chemical change
 40. physical change
 41. In a solid, particles are packed together in relatively fixed positions. The particles vibrate about a fixed point. In a liquid, particles are close together but can flow around one another. In a gas, particles are at great distances from one another, compared to the particles of liquids and solids.
 42. The composition of a pure substance is the same throughout and does not vary from sample to sample. Pure substances include elements and compounds. A mixture contains more than one substance and can vary in composition from one sample to another and sometimes within different parts of the same sample.
 43. Although energy can be absorbed or released in a

change, it is not destroyed or created. It simply takes a different form.

43. Homogeneous mixtures are uniform in composition; heterogeneous mixtures are not.
44. An element cannot be broken down, but water can be decomposed into hydrogen and oxygen by passing an electrical current through it (electrolysis).

Measurements and Calculations, pp. 5-8

- | | |
|------|------|
| a | 5. c |
| 2. c | 6. c |
| 3. c | 7. a |
| 4. a | 8. a |
9. time
 10. mass
 11. density
 12. energy
 13. length or distance
 14. volume
 15. area
 16. qualitative
 17. quantitative
 18. qualitative
 19. quantitative
 20. 3.00×10^7 km/s
 21. three
 22. 0.026 g or 2.6×10^{-2} g
 23. 2.5×10^{-4} L
 24. quantity
 25. derived units
 26. conversion factor
 27. 0.0432 kg or 4.32×10^{-2} kg
 28. 5400 mL or 5.4×10^3 mL
 29. 300. K
 30. 1.05×10^3 J
 31. 3.51×10^3 cm
 32. A hypothesis is a testable statement that can be used to make predictions and to carry out further experiments. A theory is a broad generalization that explains a body of facts or phenomena.
 33. Mass is a measure of the quantity of matter. Weight is a measure of the gravitational pull on matter. Mass does not depend on gravitational attraction.
 34. Models are visual, verbal, or mathematical representations. They are used to explain how phenomena occur or how data or events are related.
 35. The graph of two quantities that are directly proportional is a straight line through the origin.

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Jianjun Gao



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Children's Books in Print, 2007, 2006 **Subject Guide to Books in Print**, 1991 **Supplemental Problems - Glencoe Chemistry Matter and Change**, 2013 *Glencoe Chemistry* Thandi Buthelezi, Dinah Zike, 2017 **Glencoe Chemistry**, 2012 **Glencoe Chemistry Standardized Test Practice** McGraw-Hill/Glencoe, 2007-05 Chemistry Laurel Dingrando, McGraw-Hill Staff, Glencoe/McGraw-Hill, 2001-05-01 **Glencoe Chemistry** Glencoe/McGraw-Hill, 2002

Chemistry: Matter & Change, Study Guide For Content Mastery, Student Edition McGraw Hill, 2001-03-30. Study Guide and Reinforcement Worksheets allow for differentiated instruction through a wide range of question formats. There are worksheets and study tools for each section of the text that help teachers track students' progress toward understanding concepts. Guided Reading Activities help students identify and comprehend the important information in each chapter.

Chemistry, 2013 **Chemistry: Matter & Change, Laboratory Manual, Student Edition** McGraw Hill, 2012-03-12. The student edition Laboratory Manual follows the chapter sequence and reinforces concepts taught in *Chemistry Matter and Change* but can be used with any chemistry text. Students record data and conclusions directly on lab worksheets.

Glencoe Chemistry: Matter & Change, Science Notebook, Student Edition McGraw Hill, 2012-03-05. Based on the Cornell note-taking format, this resource incorporates writing into the learning process. Directly linked to the student text, this notebook provides a systematic approach to learning science by encouraging students to engage by summarizing and synthesizing abstract concepts in their own words. **Chemistry** Laurel Dingrando, McGraw-Hill Staff, Glencoe/McGraw-Hill, 2001-08-01

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