

Question 2

- Which of the following salts has the least solubility in water at 25°C?

F CaCO_3

G KBr

H CaCl_2

J KI

Ion	General Solubility Rule
NO_3^-	All nitrates are soluble
$\text{C}_2\text{H}_3\text{O}_2^-$	All acetates are soluble ($\text{AgC}_2\text{H}_3\text{O}_2$ only moderately)
Cl^- , Br^- , I^-	All chlorides, bromides and iodides are soluble except Ag^+ , Pb^+ and Hg_2^{2+} . (PbCl_2 is slightly soluble in cold water and moderately soluble in hot water.)
SO_4^{2-}	All sulfates are soluble except those of Ba^{2+} , Pb^{2+} , Ca^{2+} and Sr^{2+}
CO_3^{2-} and PO_4^{3-}	All carbonates and phosphates are insoluble except those of Na^+ , K^+ and NH_4^+ . (Many acid phosphates are soluble).
OH^-	All hydroxides are insoluble except those of Na^+ and K^+ . Hydroxides of Ba^{2+} and Ca^{2+} are slightly soluble.
S^{2-}	All sulfides are insoluble except those of Na^+ , K^+ , NH_4^+ and those of the alkaline earths: Mg^{2+} , Ca^{2+} , Sr^{2+} and Ba^{2+} . (Sulfides of Al^{3+} and Cr^{3+} hydrolyze and precipitate as the corresponding hydroxides.)
Na^+ , K^+ and NH_4^+	All salts of sodium ion, potassium ion and ammonium ion are soluble except several uncommon ones.

Lab Solubility Datasheet Answer Key

S Ben Porath



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Chemical Interactions ,2005 **Video Source Book** ,2008

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Table of Contents Lab Solubility Datasheet Answer Key

1. Understanding the eBook Lab Solubility Datasheet Answer Key
 - The Rise of Digital Reading Lab Solubility Datasheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Solubility Datasheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Solubility Datasheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Solubility Datasheet Answer Key
 - Personalized Recommendations
 - Lab Solubility Datasheet Answer Key User Reviews and Ratings
 - Lab Solubility Datasheet Answer Key and Bestseller Lists
5. Accessing Lab Solubility Datasheet Answer Key Free and Paid eBooks
 - Lab Solubility Datasheet Answer Key Public Domain eBooks
 - Lab Solubility Datasheet Answer Key eBook Subscription Services
 - Lab Solubility Datasheet Answer Key Budget-Friendly Options
6. Navigating Lab Solubility Datasheet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Solubility Datasheet Answer Key Compatibility with Devices
 - Lab Solubility Datasheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Solubility Datasheet Answer Key
 - Highlighting and Note-Taking Lab Solubility Datasheet Answer Key
 - Interactive Elements Lab Solubility Datasheet Answer Key
8. Staying Engaged with Lab Solubility Datasheet Answer Key

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Solubility Datasheet Answer Key
9. Balancing eBooks and Physical Books Lab Solubility Datasheet Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Solubility Datasheet Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lab Solubility Datasheet Answer Key
- Setting Reading Goals Lab Solubility Datasheet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lab Solubility Datasheet Answer Key
- Fact-Checking eBook Content of Lab Solubility Datasheet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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