

CHEMICAL REACTIONS AND EQUATIONS

VERY SHORT ANSWER TYPE QUESTIONS [1 MARK]

1. In electrolysis of water, why is the volume of gas collected over one electrode double that of gas collected over the other electrode?

Answer. It is because water contains hydrogen and oxygen in the ratio of 2 : 1.

2. Balance the following chemical equations.

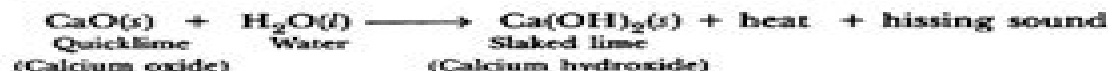


Answer.



3. What happens chemically when quicklime is added to water filled in a bucket?

Answer. Quicklime reacts with water to form slaked lime and produces lot of heat and hissing sound.



4. On what basis is a chemical equation balanced?

Answer. A chemical reaction is balanced on the basis of law of conservation of mass.

5. What change in colour is observed when white silver chloride is left exposed to sunlight? State the type of chemical reaction in this change.

Answer. Silver chloride becomes grey. It is a photochemical decomposition reaction.

6. Write a balanced chemical equation for the reaction between sodium chloride and silver nitrate indicating the physical state of the reactants and the products.

Answer.



7. State one basic difference between a physical change and a chemical change.

Answer. In physical change, no new substance is formed, whereas in a chemical change, new substance(s) is/are formed.

8. What is meant by a chemical reaction?

Answer. The reaction representing a chemical change is called a chemical reaction.

9. $\text{AgNO}_3(aq) + \text{NaCl}(aq) \xrightarrow{\hspace{2cm}} \text{AgCl}(s) \downarrow + \text{NaNO}_3(aq)$



Consider the above mentioned two chemical equations with two different kinds of arrows ($\uparrow$ and $\downarrow$) along with product. What do these two different arrows indicate?

Answer. $\uparrow$ shows the gas is evolved whereas $\downarrow$ shows insoluble substance (precipitate) is formed.

10. Hydrogen being a highly inflammable gas and oxygen being a supporter of combustion, yet water which is a compound made up of hydrogen and oxygen is used to extinguish fire. Why?

Answer. It is because properties of compound (H_2O) are different from properties of its constituting elements, i.e. H_2 and O_2 .

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