

# Gas Laws Problem Packet

**Gas Laws Packet**

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

**Essentials**

Standard Constant:

Unit	Standard Pressure
atm	1.013
mmHg	760
cmHg	76
torr	760
atm	1.013

**22.4 L = 1 mol at STP**

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law:	$P_1V_1 = P_2V_2$
Charles's Law:	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law:	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Avogadro's Law:	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law:	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Graham's Law:	$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$

The constant R (used in the ideal gas law) depends on the unit of pressure:

- 1 atm,  $R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$
- 1 atm,  $R = 8.314 \frac{\text{J}}{\text{mol}\cdot\text{K}}$

**Always Remember:**

- Temperature must always be in Kelvin! ( $K = ^\circ C + 273$ )
- If you are given two different units of pressure, CONVERT!
- Create two columns before solving: initial and final. Write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (For Ideal Gas Law, write what you're given, insert those in equation and solve for unknown.)
- Avogadro's Law: DON'T FORGET TO USE MOLE! (MAYBE)
- Graham's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation! (This is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things!

**Charles' Law  $V \propto T$**

Directions: Complete each of the problems below. Show how you set up the problem and solve.

- The temperature inside my refrigerator is about 4°C. If I place a balloon in the refrigerator, will the volume of air in the balloon increase or decrease? Explain your answer. (15 pts)
- A balloon is inflated to a volume of 2.0 L at 25°C. If the temperature is lowered to 10°C, what will the volume of the balloon be? (15 pts)

**Gay-Lussac's Law  $P \propto T$**

- A gas is held at a constant volume of 2.0 L and a pressure of 1.0 atm. If the temperature is increased to 2.0 times the original temperature, what will the pressure be? (15 pts)
- A gas is held at a constant volume of 2.0 L and a pressure of 1.0 atm. If the temperature is decreased to 0.5 times the original temperature, what will the pressure be? (15 pts)

**Avogadro's Law  $V \propto n$**

- A gas is held at a constant pressure and temperature. If the volume is increased to 2.0 times the original volume, what will the number of moles be? (15 pts)
- A gas is held at a constant pressure and temperature. If the number of moles is increased to 2.0 times the original number of moles, what will the volume be? (15 pts)

**Combined Gas Law  $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$**

- A gas is held at a constant pressure and temperature. If the volume is increased to 2.0 times the original volume, what will the number of moles be? (15 pts)
- A gas is held at a constant pressure and temperature. If the number of moles is increased to 2.0 times the original number of moles, what will the volume be? (15 pts)

**Graham's Law  $\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$**

- Gas A effuses at a rate that is 1.5 times that of Gas B. What is the ratio of the molecular masses of Gas A to Gas B? (15 pts)
- Gas A effuses at a rate that is 1.5 times that of Gas B. What is the ratio of the molecular masses of Gas B to Gas A? (15 pts)

All Gas Laws Included!



# Gas Law Problems Packet

**D Keegan**



**Gas Law Problems Packet:**

## **Gas Law Problems Packet** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Gas Law Problems Packet**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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