

- 7) If I have 21 moles of gas held at a pressure of 78 atm and a temperature of 1000 K, what is the volume of the gas?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{21 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 1000 \text{ K}}{78 \text{ atm}}$$

$$V = 17.97 \text{ L}$$

- 8) If I have 1.9 moles of gas held at a pressure of 3 atm and in a container with a volume of 50 liters, what is the temperature of the gas?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{3 \text{ atm} \cdot 50 \text{ L}}{1.9 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}}$$

$$T = 1602.07 \text{ K}$$

- 9) If I have 2.4 moles of gas held at a temperature of 47 °C and in a container with a volume of 47 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{2.4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 320 \text{ K}}{47 \text{ L}}$$

$$P = 1.62 \text{ atm} \quad P = 169.64 \text{ kPa}$$

- 10) If I have an unknown quantity of gas held at a temperature of 1100 K in a container with a volume of 27 liters and a pressure of 780 atm, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{780 \text{ atm} \cdot 27 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 1100 \text{ K}}$$

$$n = 0.0076009 \text{ mol}$$

- 11) If I have 6.275 moles of gas at a temperature of 77 K and a pressure of 1.75 atmospheres, what is the volume of the gas?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{6.275 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 77 \text{ K}}{1.75 \text{ atm}}$$

$$V = 0.976 \text{ L}$$

- 12) If I have 72 liters of gas held at a pressure of 3.4 atm and a temperature of 225 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{3.4 \text{ atm} \cdot 72 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 225 \text{ K}}$$

$$n = 0.0796 \text{ mol}$$

# Gas Laws Section 2 Review Answers

**Ralph Thomas Myers**



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