

KEY

Gas Stoichiometry Worksheet

Directions: Use the gas laws we have learned to solve each of the following problems. Each of the chemical equations must first be balanced. Show all your work for credit.

1. When calcium carbonate is heated strongly, carbon dioxide gas is released according to the following equation:



What volume of $\text{CO}_2(g)$, measured at STP, is produced if 15.2 grams of $\text{CaCO}_3(s)$ is heated?

$$\frac{15.2 \text{ g CaCO}_3}{100.1 \text{ g}} \times \frac{1 \text{ mol CaCO}_3}{1 \text{ mol CaCO}_3} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol CaCO}_3} \times \frac{22.4 \text{ L CO}_2}{1 \text{ mol CO}_2} = 3.40 \text{ L CO}_2$$

2. The synthesis of sodium chloride occurs according to the reaction:



What volume of chlorine at STP is necessary for the complete reaction of 4.81 grams of sodium metal?

$$\frac{4.81 \text{ g Na}}{23.0 \text{ g}} \times \frac{1 \text{ mol Na}}{1 \text{ mol Na}} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \times \frac{22.4 \text{ L Cl}_2}{1 \text{ mol Cl}_2} = 2.34 \text{ L Cl}_2$$

3. Potassium permanganate is produced commercially by the this reaction.



What volume of chlorine gas at STP would be required to produce 10.0 grams of KMnO_4 ?

$$\frac{10.0 \text{ g}}{158.0 \text{ g}} \times \frac{1 \text{ mol KMnO}_4}{1 \text{ mol KMnO}_4} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol KMnO}_4} \times \frac{22.4 \text{ L}}{1 \text{ mol}} = 0.709 \text{ L Cl}_2$$

Gas Stoichiometry Problems And Answers

Kaplan Test Prep



Gas Stoichiometry Problems And Answers:

This book delves into Gas Stoichiometry Problems And Answers. Gas Stoichiometry Problems And Answers is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Gas Stoichiometry Problems And Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Gas Stoichiometry Problems And Answers
- Chapter 2: Essential Elements of Gas Stoichiometry Problems And Answers
- Chapter 3: Gas Stoichiometry Problems And Answers in Everyday Life
- Chapter 4: Gas Stoichiometry Problems And Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Gas Stoichiometry Problems And Answers. This chapter will explore what Gas Stoichiometry Problems And Answers is, why Gas Stoichiometry Problems And Answers is vital, and how to effectively learn about Gas Stoichiometry Problems And Answers.

3. In chapter 2, the author will delve into the foundational concepts of Gas Stoichiometry Problems And Answers. The second chapter will elucidate the essential principles that must be understood to grasp Gas Stoichiometry Problems And Answers in its entirety.

4. In chapter 3, the author will examine the practical applications of Gas Stoichiometry Problems And Answers in daily life. The third chapter will showcase real-world examples of how Gas Stoichiometry Problems And Answers can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of Gas Stoichiometry Problems And Answers in specific contexts. This chapter will explore how Gas Stoichiometry Problems And Answers is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Gas Stoichiometry Problems And Answers. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Gas Stoichiometry Problems And Answers.

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