

Format for Final Report

The final report is a comprehensive accounting of the three functional areas – administrative, education/outreach, and scientific/research – funded under the award. The NOAA Program Officer will review and provide feedback on acceptance of the Final Report. The NOAA Grants Management Division reviews the financial reporting and working with the Program Office determines if the full Final Report is approved. Award closeout is initiated after all reports are approved by NOAA.

Your CSC final report must stand on its own, such that a Program Officer need not refer to your proposal in order to understand the relationship between your report and the funded project. Effective reports include a succinct statement describing the originally proposed project (hypotheses, goals, strategies, and methods) at the beginning of the Approach and the section. Effective reports make links between the original hypotheses, goals, and/or strategies described in your proposal, approved plans, and the various findings and completed activities from the performing period. The content of your report, including PDF attachments, should not depend on access to internet hyperlinks. Please write a report that is as to the point, complete, and accurate as possible.

Guidance for Final Report

I. Report Title, Authors, Institutions/Organizations, Award Number, Award Performing Period (Start Date through End Date)

II. CSC Abstract

A brief (one paragraph) description of the Final Report.

III. CSC Project Outcomes Report

A. A stand-alone description of the CSC award that is intended for viewing by the general public, to present a complete picture of the results of the funded CSC.

Note: There is increasing public and congressional interest in assuring that federal dollars are well spent, so while the outcomes report should be brief and accurate, PIs should strive to help the public appreciate the value of the work completed. The Outcomes Report should:

- not include any unpublished information, data, inventions or any other confidential information.
- not include any personally identifiable information such as names of participants.

Format For Final Report

Raffaela Di Napoli



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clayfield rita returned to god on 22 may 1457 and her body is venerated in the shrine

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
one mole of HCl would be fully neutralized by one mole of NaOH

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{NaOH} \cdot \text{Na}_2\text{C}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction
NaOH + CH₃COOH → NaCH₃COO + H₂O
the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of CH₃COOH in the

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