

## Density Worksheet

### Physical Science

$D = m/V$

Densities of Common Substances @ 20°C				
Substance	Density (g/cm <sup>3</sup> )		Substance	Density (g/cm <sup>3</sup> )
Oxygen	0.00133		Aluminum	2.70
Hydrogen	0.000084		Iron	7.87
Ethanol	0.785		Copper	8.96
Benzene	0.880		Silver	10.5
Water	1.000		Lead	11.34
Magnesium	1.74		Mercury	13.6
Salt (sodium chloride)	2.16		Gold	19.32

1. The ratio of an object's mass to its \_\_\_\_\_ is called the *density* of the object.
  
2. A kilogram of lead occupies a much smaller volume than a kilogram of water, because \_\_\_\_\_ has a much higher *density*.
  
3. For the masses and volumes indicated, calculate the **density** in grams per cubic centimeters.
  - A. mass = 453 g; volume = 225 cm<sup>3</sup>
  - B. mass = 5.0 g; volume = 10.0 cm<sup>3</sup>
  - C. mass = 26.1 g; volume = 2.0 mL
  
4. If 89.2 mL of a liquid has a mass of 75.2 g, calculate the liquid's density.
  
  
5. A cube of metal weighs 1450 g and displaces 542 mL of water when immersed. Calculate the density of the metal.

# Everyday Physical Science Answers

**F Rizvi**



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## **Table of Contents Everyday Physical Science Answers**

1. Understanding the eBook Everyday Physical Science Answers
  - The Rise of Digital Reading Everyday Physical Science Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Everyday Physical Science Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Everyday Physical Science Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Everyday Physical Science Answers
  - Personalized Recommendations
  - Everyday Physical Science Answers User Reviews and Ratings
  - Everyday Physical Science Answers and Bestseller Lists
5. Accessing Everyday Physical Science Answers Free and Paid eBooks
  - Everyday Physical Science Answers Public Domain eBooks
  - Everyday Physical Science Answers eBook Subscription Services
  - Everyday Physical Science Answers Budget-Friendly Options
6. Navigating Everyday Physical Science Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Everyday Physical Science Answers Compatibility with Devices
  - Everyday Physical Science Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Everyday Physical Science Answers
  - Highlighting and Note-Taking Everyday Physical Science Answers
  - Interactive Elements Everyday Physical Science Answers
8. Staying Engaged with Everyday Physical Science Answers

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Everyday Physical Science Answers
9. Balancing eBooks and Physical Books Everyday Physical Science Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Everyday Physical Science Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Everyday Physical Science Answers
- Setting Reading Goals Everyday Physical Science Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Everyday Physical Science Answers
- Fact-Checking eBook Content of Everyday Physical Science Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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