

## Chem 162A

### Chapter 14. Covalent Bonding: Orbitals

#### → Hybridization and the LE Model of Bonding

- Lewis structures of molecules
- prediction of geometry of molecules
- hybrid orbitals ( $sp^3$ ,  $sp^2$ ,  $sp$ ,  $dsp^2$ ,  $d^2sp^3$ )
- interpretation of structure and bonding

#### → Molecular Orbital Model of Bonding in Molecules

- molecular orbital diagrams
- bond order
- magnetism

#### → Molecular Spectroscopy

- Electronic Spectroscopy
- Vibrational/Rotational Spectroscopies
- Nuclear Magnetic Resonance (NMR) Spectroscopy

#### 14.1 Hybridization and the LE Model of Bonding

- Atomic bonding involves only valence orbitals
- Methane,  $CH_4$



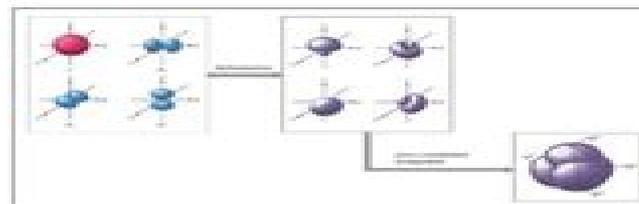
- H atoms in  $CH_4$  will use 1s orbitals
- Of the two types of orbitals (2s and 2p) which will C atoms use for bonding in  $CH_4$ ?
  - If both are used: 2 different types of C-H bonds (Contrary to experimental facts)
  - Neither of the "native" atomic orbitals of C atoms are used, instead, new hybrid orbitals are used.

#### Hybridization of atomic orbitals

- The mixing of the "native" atomic orbitals to form special orbitals for bonding is called **hybridization**.
- The 4 new **equivalent** orbitals formed by mixing the one 2s and three 2p orbitals are called  $sp^3$  orbitals.
- The carbon atom is said to undergo  $sp^3$  hybridization, i.e. is  $sp^3$  hybridized.
- Energy-level diagram showing the  $sp^3$  hybridization.



Figure 14.3: native 2s and three 2p atomic orbitals characteristic of a free carbon atom are combined to form a new set of four  $sp^3$  orbitals.



# General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide

**Yi-Tong Ma**



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