

DAT General Chemistry Equation Sheet

Chapter 0: General and Lab Concepts Review

Dilutions	$M_1V_1 = M_2V_2$ or $C_1V_1 = C_2V_2$	M or C = concentration V = volume
Percent Error	$\frac{(A - T)}{T} \times 100$	T = theoretical A = actual
Absorbance (Spectrophotometer)	$Abs = \epsilon cl$	ϵ = molar extinction coefficient (molar absorptivity) c = sample's concentration l = path length

Chapter 2: Atomic and Electronic Structure

Energy of a photon	$E_{\text{photon}} = hf = \frac{hc}{\lambda}$	h = Planck's constant ($6.63 \times 10^{-34} \text{ J} \cdot \text{s}$) f = photon's frequency c = speed of light ($3.0 \times 10^8 \text{ m/s}$) λ = photon's wavelength
Absorption/Emission Line Spectra	$\Delta E = E_{\text{photon}}$	
Kinetic Energy of an electron (Photoelectric Effect)	$KE_{e^-} = E_{\text{photon}} - \phi$	ϕ = work function (minimum energy needed to ionize electron)

Chapter 7: Chemical Solutions

Molarity	$M = \frac{\text{moles}_{\text{solute}}}{L_{\text{solution}}}$	
Molality	$m = \frac{\text{moles}_{\text{solute}}}{kg_{\text{solvent}}}$	
Henry's Law	$P_A = k_H[A]$	P_A = partial pressure of gas A k_H = Henry's Law constant (varies per problem) $[A]$ = conc. of gas A
Freezing Point Depression	$\Delta T_f = -iK_f m$	i = van't Hoff factor K_f = F.P. depression constant m = molality
Boiling Point Elevation	$\Delta T_B = iK_B m$	i = van't Hoff factor K_B = B.P. elevation constant m = molality
Vapor Pressure Depression (Raoult's Law)	$P_{\text{solute}} = X_{\text{solute}} P_{\text{solute}}^{\circ}$	P_{solute} = VP of solution X_{solute} = mol fract of solvent $P_{\text{solute}}^{\circ}$ = VP of pure solvent
Osmotic Pressure (π)	$\pi = iMRT$	M = molarity of solute i = van't Hoff factor R = $0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$ T = temp. in Kelvin

Chapter 5: Gases

Pressure	$P = \frac{F}{A}$	F = force A = area
Average Kinetic Energy	$KE_{\text{avg}} = \frac{3}{2} RT$	$R = 0.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$
Root-Mean-Square Speed (v_r)	$v_r = \sqrt{\frac{3RT}{M_m}}$	$R = 0.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$ M_m = molar mass
Ideal Gas Law	$PV = nRT$	n = # of moles $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$
Boyle's Law	$V \propto \frac{1}{P}$	
Charles' Law	$V \propto T$	
Avogadro's Law	$V \propto n$	
Combined Gas Law	$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$	
Standard Temp. & Pressure (STP) Standard Conditions	$P = 1 \text{ atm}$ $T = 273 \text{ K}$ All aqueous species @ 1 M All gaseous species @ 1 atm	*1 mol of gas = 22.4 L at STP
Density	$\frac{P(MM)}{RT} = \frac{m}{V}$	MM = molar mass $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$ m = mass V = volume
Dalton's Law of Partial Pressures	$P_{\text{total}} = P_A + P_B + \dots$	
Dalton's Law of Partial Pressures	$P_A = X_A P_{\text{total}}$	X_A = mol fraction of gas A
Graham's Law of Effusion	$\frac{r_1}{r_2} = \sqrt{\frac{M_{m2}}{M_{m1}}}$	r = rate of effusion M = molar mass
Real Gas Equation	$(P + \frac{an^2}{V^2})(V - nb) = nRT$	a & b = constants specific to each gas $\frac{an^2}{V^2}$ corrects for IMF's $-nb$ corrects for volume

Chapter 8: Chemical Kinetics

General	$A + B \rightarrow C + D$	k = rate constant m & n = determined experimentally
Rate Law	$\text{rate} = k[A]^m[B]^n$	k = rate constant M = molarity t = seconds
Rate Constant Units	0 order: $k = M^{-1} \cdot s^{-1}$ 1 st order: $k = s^{-1}$ 2 nd order: $k = M^{-2} \cdot s^{-1}$ 3 rd order: $k = M^{-3} \cdot s^{-1}$	
Arrhenius Equation	$k = Ae^{-E_a/RT}$	k = rate constant A = unique to each rxn E_a = act. energy $R = 0.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$ T = temp. in Kelvin

General Chemistry Formula And Equations Sheet

Joseph William Mellor



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