

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table



LIMITING REACTANTS

1 The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you start with 28 grams of N_2 and 12 grams of H_2 , determine the limiting reactant.

2 Calculate the grams of ammonia produced.

3 How many grams of the excess reactant is present in a left over?
* Multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

4 Calculate moles of product * mole of product.

5 The reaction below produces carbon dioxide and water from a common household product.
 $2C_2H_5OH + O_2 \rightarrow 2CO_2 + 2H_2O$
If you have 100 grams of ethyl alcohol and 50 grams of oxygen, determine the limiting reactant.

6 How many grams of water produced are produced?



7 Magnesium chloride is used to create salts for deicing roads. The reaction with very impure iron is produced.
 $Mg + 2FeCl_3 \rightarrow MgCl_2 + 2Fe$
If you start with 20 grams of magnesium and 2 moles of $FeCl_3$, determine the limiting reactant.

8 Calculate the grams of magnesium chloride produced.

9 How many grams of the excess reactant is present in a left over?

10 The reaction below is used to create a fuel for heating and cooking.
 $2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O$
If you have 200 grams of propane and 50 grams of oxygen, determine the limiting reactant.

11 How much carbon dioxide is produced?

12 The reaction below produces ammonia for use in fertilizers and a common household product.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you have 100 grams of nitrogen and 50 grams of hydrogen, determine the limiting reactant.

13 Calculate the grams of ammonia produced.

14 How many grams of the excess reactant is present in a left over?

15 Calculate the grams of carbon dioxide produced.

16 How many grams of the excess reactant is present in a left over?

ANSWER KEY INCLUDED

Limiting Reactant Answer Key

**Spencer L. Seager, Michael R.
Slabaugh**

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