

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table



LIMITING REACTANTS

1. The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you start with 28 grams of N_2 and 12 grams of H_2 , determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is produced the left over?
* Multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

4. Calculate moles of product * mole of grams.

5. The reaction below produces carbon dioxide and water from a common household product. Also, carbon dioxide is used in soft drinks.
 $C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O$

If you have 100 grams of ethyl alcohol and 40 grams of oxygen, determine the limiting reactant.

6. How many grams of water produced are produced?



7. Magnesium reacts with oxygen to create white for welding rods. The reaction is:
 $Mg + O_2 \rightarrow MgO$
If you start with 24 grams of magnesium and 16 grams of oxygen, determine the limiting reactant.

8. Calculate the grams of magnesium oxide produced.

9. How many grams of the excess reactant is produced the left over?

10. The reaction below is used to create carbon dioxide for soft drinks.
 $C_2H_4O_2 + 2O_2 \rightarrow 2CO_2 + 2H_2O$
If you have 100 grams of acetic acid and 40 grams of oxygen, determine the limiting reactant.

11. How many carbon dioxide is produced?

12. The reaction below is used to create ammonia for use in fertilizers and household cleaning products.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you have 28 grams of nitrogen and 12 grams of hydrogen, determine the limiting reactant.

13. Calculate the grams of ammonia produced.

14. How many grams of the excess reactant is produced the left over?

15. Calculate the grams of the excess reactant is produced the left over?

ANSWER KEY INCLUDED

Limiting Reactant Problems Answer Key

James E. Brady, Fred Senese



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Prentice Hall Chemistry, 2000

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