

For the following reactions, find the following:

- a) Which of the reagents is the limiting reagent?
- b) What is the maximum amount of each product that can be formed?
- c) How much of the other reagent is left over after the reaction is complete?

1) Consider the following reaction:



Answer the questions above, assuming we started with 30 grams of ammonium nitrate and 50 grams of sodium phosphate.

30.0 g NH ₄ NO ₃	1 mole	3 moles NaNO ₃	85.00 g NaNO ₃	-31.9 g NaNO ₃ <i>max amount!</i>
80.06 g	3 moles NH ₄ NO ₃	1 mole NaNO ₃	85.00 g NaNO ₃	

limiting reactant is NH₄NO₃

50 g Na ₃ PO ₄	1 mole Na ₃ PO ₄	3 moles NaNO ₃	85.00 g NaNO ₃	= 77.8 g NaNO ₃
163.94 g	1 mole Na ₃ PO ₄	1 mole NaNO ₃	1 mole NaNO ₃	

ⓐ 30.0 g NH₄NO₃

1 mole NH ₄ NO ₃	1 mole Na ₃ PO ₄	163.94 g	20.5 g Na ₃ PO ₄ needed
80.06 g NH ₄ NO ₃	3 moles NH ₄ NO ₃	1 mole Na ₃ PO ₄	

left over: 50.0 - 20.5 = 29.5 g left over
start used

Limiting Reagent Test 2014 Form Answers

**Theodore L. Brown, H. Eugene LeMay
Jr., Bruce E. Bursten, Catherine
Murphy, Patrick Woodward, Steven
Langford, Dalius Sagatys, Adrian
George**

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