

## Student Exploration: Limiting Reactants

**Essential Question:** How do you determine which reactant is the limiting reactant in a chemical reaction?

**What You Will Do:** You will use the Gizmo to explore how limiting reactants affect the amount of product that can be formed in a chemical reaction.

1. How many molecules of product can you make? \_\_\_\_\_ Molecules of product

2. Which reactant completely runs out first? \_\_\_\_\_ The limiting reactant is \_\_\_\_\_.

3. Which reactant will you have left over? \_\_\_\_\_ The excess reactant is \_\_\_\_\_.

### What's Going On?

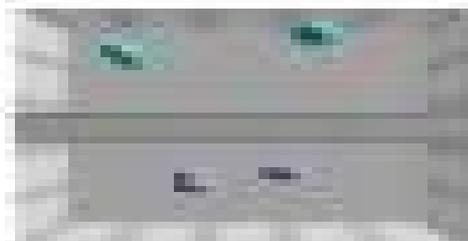
Just as ingredients can be used together to make a limited number of sandwiches, a limited number of molecules can be used to make a limited number of products. The reactant that runs out first is the **limiting reactant**. The reactant that is left over is the **excess reactant**.

Remember, always use the correct units. You have information on the number of molecules of each reactant. The number of molecules of product is always the same as the number of limiting reactant molecules.

The reactants are  $\text{H}_2$  and  $\text{O}_2$ . The product is  $\text{H}_2\text{O}$ . The ratio of the  $\text{H}_2$ ,  $\text{O}_2$ , and  $\text{H}_2\text{O}$  is 2:1:2. This means that for every 2 molecules of  $\text{H}_2$ , you need 1 molecule of  $\text{O}_2$  to make 2 molecules of  $\text{H}_2\text{O}$ .

4. How many molecules of  $\text{H}_2$  and  $\text{O}_2$  do you have? \_\_\_\_\_

a. How many molecules of  $\text{H}_2$  are there? \_\_\_\_\_ The bigger number? \_\_\_\_\_



# Limiting Reactant Student Exploration Sheet Answer

**LP Steffe**



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