

Quiz & Worksheet - Limiting Reactants & Excess Reactants

1. What is a limiting reactant?

- The reactant that is used up first and prevents more product from being made.
- The reactant that is used up last and prevents more product from being made.
- The reactant that makes the product.
- The substance that is in excess that doesn't get used up as a reactant.

2. Say you take a reactant A and calculate the amount of moles of another reactant B required to use up all of A. How do you know which of two reactants is the limiting one?

- You compare the calculated amount of B to the actual amount available. If more is required, then B is the limiting reactant. If less is required, then A is the limiting reactant.
- You compare the calculated amount of B to the actual amount available. If more is required, then A is the limiting reactant. If less is required, then B is the limiting reactant.
- If the calculated value of B is larger than the amount of A, then A is the limiting reactant.
- If the calculated value of B is larger than the amount of A, then B is the limiting reactant.

3. If 2.5 moles of zinc react with 6.0 moles of hydrochloric acid in the equation $\text{Zn} + 2\text{HCl} = \text{ZnCl}_2 + \text{H}_2$, what is the limiting reactant?

- Zn
- HCl
- ZnCl
- ZnH

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Multicultural applications of mastery learning Güzver Yıldırım,2006

Limiting Reactant and Percent Lifeliqe,2019

This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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