

Nuclear Chemistry - Guided Practice Questions**Multiple Choice** - Choose the one best answer and write its letter in the blank provided.

- ____ 1) In balancing the nuclear reaction ${}^{238}_{92}\text{U} \rightarrow {}^{234}_{90}\text{Th} + \text{?He}$, the identity of the missing element is _____?
 A) Pu B) U C) Th D) Np E) Pa
- ____ 2) The following rxn, ${}^{210}_{84}\text{Po} \rightarrow {}^{210}_{82}\text{Pb} + \text{_____}$, is an example of what form of radioactive decay?
 A) alpha B) positron C) gamma D) beta E) e-capture
- ____ 3) The missing product from this reaction is: ${}^{238}_{92}\text{U} \rightarrow {}^{234}_{90}\text{Th} + \text{_____?}$
 A) ${}^4_2\text{He}$ B) ${}^0_{-1}\text{e}$ C) ${}^4_2\text{He}$ D) ${}^0_1\text{n}$ E) ${}^0_{-1}\text{e}$
- ____ 4) What is the mass number of a neutron _____?
 A) 1 B) 3 C) 2 D) 0 E) 4
- ____ 5) In the nuclear decay reaction represented by ${}^{238}_{92}\text{U} \rightarrow {}^{234}_{90}\text{Th}$, the emitted particle is _____?
 A) beta particle B) neutron C) alpha particle D) positron E) proton
- ____ 6) The reaction shown below is responsible for creating C-14 in the atmosphere. What is the bombarding particle _____?
 ${}^{14}_7\text{N} + \text{_____} \rightarrow {}^{14}_6\text{C} + \text{?H}$
 A) electron B) proton C) positron D) alpha particle E) neutron
- ____ 7) ${}^{131}_{53}\text{I}$ has a half-life of 8.04 days. Assuming you start with a 1.53 mg sample of ${}^{131}_{53}\text{I}$, how many mg will remain after 13.0 days _____?
 A) 0.440 B) 0.422 C) 0.298 D) 0.835 E) 0.499
- ____ 8) All atoms of a given element have the same:
 A) atomic mass, B) mass number, C) # of nucleons, D) atomic #, E) # of neutrons.
- ____ 9) Atoms containing radioactive nuclei are called:
 A) radioisotopes, B) radionuclides, C) nuclides, D) nucleons, E) radioisophores.
- ____ 10) What happens to the mass number and the atomic number of an element when it undergoes beta decay?
 A) The mass number decreases by 4 and the atomic number decreases by 2.
 B) The mass number does not change and the atomic number decreases by 2.
 C) Neither the mass number nor the atomic number change.
 D) The mass number increases by 2 and the atomic number increases by 1.
 E) The mass number does not change and the atomic number increases by 1.

Guided Practice Problem Chemistry

Karen Timberlake



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Table of Contents Guided Practice Problem Chemistry

1. Understanding the eBook Guided Practice Problem Chemistry
 - The Rise of Digital Reading Guided Practice Problem Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Practice Problem Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Practice Problem Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Practice Problem Chemistry
 - Personalized Recommendations
 - Guided Practice Problem Chemistry User Reviews and Ratings
 - Guided Practice Problem Chemistry and Bestseller Lists

5. Accessing Guided Practice Problem Chemistry Free and Paid eBooks
 - Guided Practice Problem Chemistry Public Domain eBooks
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 - Guided Practice Problem Chemistry Budget-Friendly Options
6. Navigating Guided Practice Problem Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Guided Practice Problem Chemistry Compatibility with Devices
 - Guided Practice Problem Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guided Practice Problem Chemistry
 - Highlighting and Note-Taking Guided Practice Problem Chemistry
 - Interactive Elements Guided Practice Problem Chemistry
8. Staying Engaged with Guided Practice Problem Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Practice Problem Chemistry
9. Balancing eBooks and Physical Books Guided Practice Problem Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Practice Problem Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Guided Practice Problem Chemistry
 - Setting Reading Goals Guided Practice Problem Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Practice Problem Chemistry
 - Fact-Checking eBook Content of Guided Practice Problem Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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