

## CHEMICAL EQUILIBRIUM WORKSHEET

On the line at the left, write the letter of the description that best matches each term.

- |       |                                |  |
|-------|--------------------------------|--|
| _____ | 1. Equilibrium position        | a. used to determine if a reaction has reached equilibrium   |
| _____ | 2. Law of chemical equilibrium | b. depends on the initial concentrations of the substances in a reaction   |
| _____ | 3. Reaction quotient           | c. states that every reaction proceeds to an equilibrium state with a specific $K_{eq}$                              |
| _____ | 4. Law of mass action          | d. expresses the relative concentration of reactants and products at equilibrium in terms of an equilibrium constant |
| _____ | 5. Equilibrium constant        | e. the ratio of product concentration to reactant concentration at equilibrium                                       |

Answer each of the following in the space provided

6. What is the equilibrium expression for the equation  $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$ ?
7. What is the equilibrium expression for the equation  $NH_4Cl(s) \rightleftharpoons NH_3(g) + HCl(g)$ ?
8. What is the equilibrium expression for the equation  $As_4O_6(s) + 6C(s) \rightleftharpoons As_4(g) + 6CO(g)$ ?
9. What is the equilibrium expression for the equation  $SnO_2(s) + 2CO(g) \rightleftharpoons Sn(s) + 2CO_2(g)$ ?
10. What is the equilibrium expression for the equation  $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$ ?
11. For the reaction  $2CO(g) \rightleftharpoons C(s) + CO_2(g)$ ,  $K_{eq} = 7.7 \times 10^{-28}$ . At a particular time, the following concentrations are measured:  $[CO]=0.034$  M,  $[CO_2]=3.6 \times 10^{-17}$  M. Is this reaction at equilibrium? If not which direction will the reaction proceed?
12. For the reaction  $N_2O_4(g) \rightleftharpoons 2NO_2(g)$ ,  $K_{eq} = 0.2$ . At a particular time, the following concentrations are measured:  $[N_2O_4]=2.0$  M,  $[NO_2] = 0.2$  M. Is this reaction at equilibrium? If not which direction will the reaction proceed?
13. For the reaction  $2HCl(g) \rightleftharpoons H_2(g) + Cl_2(g)$ ,  $K_{eq} = 0.11$ . At a particular time, the following concentrations are measured:  $[HCl]=2.5$  M,  $[H_2] = 2.0$  M,  $[Cl_2]= 1.2$  M. Is this reaction at equilibrium? If not which direction will the reaction proceed?
14. At 340 °C,  $K_{eq} = 0.064$  for the reaction  $Fe_2O_3(s) + 3H_2(g) \rightleftharpoons 2Fe(s) + 3H_2O(g)$  Given that  $[H_2]=0.45$  M and  $[H_2O]=0.37$  M, find Q and predict how the reaction will proceed.

Match each statement with the appropriate letter. Each letter can be used once, more than once, or not at all.

- |       |   |                                     |
|-------|---|-------------------------------------|
| _____ | 15. The equilibrium concentration of products is much greater than that of reactants. | a. $K_{eq}$ is much greater than 1. |
| _____ | 16. The equilibrium concentration of products is much less than that of reactants     | b. $K_{eq}$ is about equal to 1.    |
| _____ | 17. There is a considerable amount of both reactants and products at equilibrium      | c. $K_{eq}$ is much less than 1.    |

# Investigating Chemical Equilibrium Answer Key

**David Kirk**



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