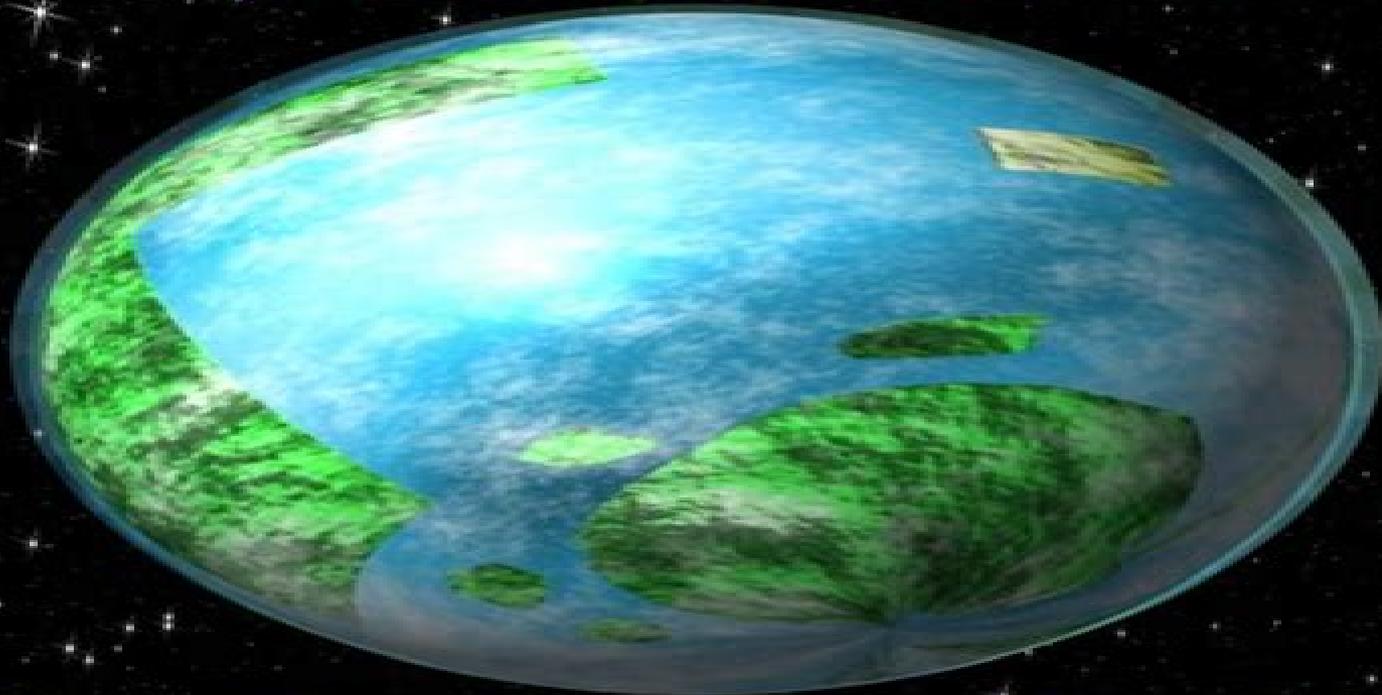


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 exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 va
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 temperature 2 the beaker now contains sodium chloride and water and the

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 measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less
 exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of
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 eg 1 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$ for which $\Delta H = 57 \text{ kJ mol}^{-1}$ eg 2 $\text{HNO}_3(\text{aq}) + \text{KOH(aq)} \rightarrow \text{KNO}_3(\text{aq}) + \text{H}_2\text{O(l)}$ for which $\Delta H = 57 \text{ kJ mol}^{-1}$

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