

Chemistry Study Guide

Periodic Table of the Elements

| | | | | | | | | | | | | | | | | | | | | | | | | | | | |
|----|----|----|----|----|----|----|----|----|----|----|----|-----|----|-----|----|-----|-----|---|----|----|----|----|----|----|----|----|----|
| 1 | 2 | | | | | | | | | | | 3 | 4 | 5 | 6 | 7 | 8 | 9 | 10 | 11 | 12 | 13 | 14 | 15 | 16 | 17 | 18 |
| H | He | | | | | | | | | | | B | C | N | O | F | Ne | | | | | | | | | | |
| Li | Be | | | | | | | | | | | Al | Si | P | S | Cl | Ar | | | | | | | | | | |
| Na | Mg | Sc | Ti | V | Cr | Mn | Fe | Co | Ni | Cu | Zn | Ga | Ge | As | Se | Br | Kr | | | | | | | | | | |
| Rb | Sr | Y | Zr | Nb | Mo | Tc | Ru | Rh | Pd | Ag | Cd | In | Sn | Sb | Te | I | Xe | | | | | | | | | | |
| Cs | Ba | | Hf | Ta | W | Re | Os | Ir | Pt | Au | Hg | Tl | Pb | Bi | Po | At | Rn | | | | | | | | | | |
| Fr | Ra | | Rf | Db | Sg | Bh | Hs | Mt | Ds | Rg | Cn | Uut | Ft | Uup | Lv | Uus | Uuo | | | | | | | | | | |
| | | La | Ce | Pr | Nd | Pm | Sm | Eu | Gd | Tb | Dy | Ho | Er | Tm | Yb | Lu | | | | | | | | | | | |
| | | Ac | Th | Pa | U | Np | Pu | Am | Cm | Bk | Cf | Es | Fm | Md | No | Lr | | | | | | | | | | | |

Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🎯

Matter:

Anything that has **mass** and **occupies space** 🌟🌍

Examples: air, gases, liquids, solidz!

Atoms:

Atoms are the **smallest, indivisible particle** 🔬👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

| Particle | Symbol | Mass (amu) | Charge | Location |
|----------|--------|------------|--------|-------------|
| Proton | p^+ | 1 | +1 | Nucleus |
| Neutron | n | 1 | 0 | Nucleus |
| Electron | e^- | ~ 0 | -1 | Orbit/Shell |

General Chemistry Study Guide

Xiaolong Qi



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Table of Contents Genral Chemistry Study Guide

1. Understanding the eBook Genral Chemistry Study Guide
 - The Rise of Digital Reading Genral Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Genral Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Genral Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Genral Chemistry Study Guide
 - Personalized Recommendations
 - Genral Chemistry Study Guide User Reviews and Ratings
 - Genral Chemistry Study Guide and Bestseller Lists
5. Accessing Genral Chemistry Study Guide Free and Paid eBooks
 - Genral Chemistry Study Guide Public Domain eBooks
 - Genral Chemistry Study Guide eBook Subscription Services
 - Genral Chemistry Study Guide Budget-Friendly Options

6. Navigating General Chemistry Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Study Guide Compatibility with Devices
 - General Chemistry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Study Guide
 - Highlighting and Note-Taking General Chemistry Study Guide
 - Interactive Elements General Chemistry Study Guide
8. Staying Engaged with General Chemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Study Guide
9. Balancing eBooks and Physical Books General Chemistry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Study Guide
 - Setting Reading Goals General Chemistry Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Study Guide
 - Fact-Checking eBook Content of General Chemistry Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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