

CHM 1102: Advanced Assignment 11A and 11B

Name: John Brown Section #: 111

Lab section #: 10

max 10

Assignment 11A

1.	Compound	Mass	Molar Mass	Molar Mass	Mass
	$\text{C}_2\text{H}_6\text{O}$	46.07	46.07	46.07	46.07
	$\text{C}_2\text{H}_4\text{O}$	44.05	44.05	44.05	44.05
	$\text{C}_2\text{H}_2\text{O}$	42.04	42.04	42.04	42.04

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Assignment 11B

1. net ionic eqn	2. net ionic eqn	3. net ionic eqn
$\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl}$	$\text{Mg}^{2+} + 2\text{OH}^- \rightarrow \text{Mg(OH)}_2$	$\text{Cu}^{2+} + \text{S}^{2-} \rightarrow \text{CuS}$

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4. net ionic eqn = $\text{Pb}^{2+} + \text{S}^{2-} \rightarrow \text{PbS}$

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net ionic eqn = $\text{Ba}^{2+} + \text{SO}_4^{2-} \rightarrow \text{BaSO}_4$

net ionic eqn = $\text{Ca}^{2+} + \text{CO}_3^{2-} \rightarrow \text{CaCO}_3$

5. net ionic eqn = $\text{Fe}^{3+} + 3\text{OH}^- \rightarrow \text{Fe(OH)}_3$

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net ionic eqn = $\text{Mg}^{2+} + 2\text{OH}^- \rightarrow \text{Mg(OH)}_2$

net ionic eqn = $\text{Ca}^{2+} + \text{CO}_3^{2-} \rightarrow \text{CaCO}_3$

6. net ionic eqn = $\text{Mg}^{2+} + \text{CO}_3^{2-} \rightarrow \text{MgCO}_3$

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net ionic eqn = $\text{Ca}^{2+} + \text{CO}_3^{2-} \rightarrow \text{CaCO}_3$

net ionic eqn = $\text{Ca}^{2+} + \text{CO}_3^{2-} \rightarrow \text{CaCO}_3$

7. net ionic eqn = $\text{Ca}^{2+} + \text{CO}_3^{2-} \rightarrow \text{CaCO}_3$

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Rosina Ehmman



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