

# Electronic Configuration

1. Use noble gases to write the abbreviated electron configurations of the following?

i. Li \_\_\_\_\_

ii. B \_\_\_\_\_

iii. O \_\_\_\_\_

iv. Mg \_\_\_\_\_

v. Cl \_\_\_\_\_

vi. V \_\_\_\_\_

vii.  $\text{Se}^{2-}$  \_\_\_\_\_

viii.  $\text{Cr}^{2+}$  \_\_\_\_\_

ix. I \_\_\_\_\_

x. Ag \_\_\_\_\_

2. Identify the element with the following electron configuration

i.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$  \_\_\_\_\_

ii.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$  \_\_\_\_\_

iii.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$  \_\_\_\_\_

iv.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$  \_\_\_\_\_

v.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^1$  \_\_\_\_\_

3. Determine if the following electron configurations are correct:

i.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^6 5s^1$  \_\_\_\_\_

ii.  $1s^2 2s^2 2p^6 3s^3$  \_\_\_\_\_

iii.  $[\text{Rn}] 7s^2 5f^9 6d^2$  \_\_\_\_\_

iv.  $[\text{Ar}] 5s^2 4d^{10} 5p^5$  \_\_\_\_\_

v.  $[\text{Xe}] 6s^2 4f^{10}$  \_\_\_\_\_

4. Identify the following elements

i.  $1s^2 2s^2 2p^2$  \_\_\_\_\_

ii.  $1s^2 2s^2 2p^6$  \_\_\_\_\_

iii.  $[\text{Ar}] 4s^2 3d^{10} 4p^5$  \_\_\_\_\_

iv.  $[\text{Kr}] 5s^2 4d^1$  \_\_\_\_\_

v.  $[\text{Ne}] 3s^2 3p^1$  \_\_\_\_\_

# Mcgraw Hill Chemistry Answer Key Electron

**Dean R. Hess, Neil R.  
MacIntyre, William F. Galvin**

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