

Factors Affecting the Rate of Chemical Reactions

What is rate of reaction and how does it apply to chemical reactions?

In a chemical reaction, how quickly or slowly reactants turn into products is called the **rate of reaction**. A reaction that takes a long time has a low reaction rate. A reaction that occurs quickly has a high reaction rate. A *rate* describes how quickly or slowly a change occurs. Every chemical reaction proceeds at a definite rate. However, you can speed up or slow down the rate of a chemical reaction.

Collision Theory

For a reaction to occur, reacting particles must collide with sufficient energy for the reacting particles to convert into the products. This relatively simple idea suggests that there are two ways we can influence the rate of the reaction. The first is increasing the number of collisions. Simply making the reacting particles collide more often should mean that there are more collisions with enough energy to make the products. The second method would be to try to increase the amount of energy in a collision. This would also mean that more collisions have enough energy to form the products. Thus collision theory suggests that increasing the number of collisions, and increasing the energy in the collisions should both increase the rate of the reaction - and vice versa.

What factors affect the rate of a chemical reaction?

The four main factors that affect the rate of chemical reactions are

- Concentration of the reactants,
- Temperature of the reactants,
- Surface area of the reactants,
- Addition of a Catalyst.

1. Effect of Concentration on reaction rate

Concentration refers to how much solute is dissolved in a solution. If a greater concentration of reactant atoms and molecules is present, there is a greater chance that collisions will occur among them. More collisions mean a higher reaction rate. Thus, increasing the concentration of the reactants usually results in a higher reaction rate. At lower concentrations, there is less chance for collisions between particles. This means that decreasing the concentrations of the reactants results in a lower reaction rate.



Low concentration = Few collisions



High concentration = More collisions

Mcgraw Hill Reaction Rates Answer Key

Michael Brown



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