

Section 5.2 Review

The Structure of the Atom

Multiple Choice

For each question below, select the letter of the best answer.

- Which of the following statements about protons, neutrons, and electrons is *true*?
 - Protons and electrons have about the same mass.
 - Neutrons and electrons have equal but opposite charges.
 - The mass of a proton is nearly 2000 times greater than the mass of an electron.
 - The mass of an proton is nearly 2000 times greater than the mass of a neutron.
 - Neutrons and electrons are found in the nucleus of an atom.
- What is the atomic number of mercury, Hg?
 - 2+
 - 80
 - 120
 - 200
 - 200.6
- What is the mass number of an atom of chlorine with 18 neutrons?
 - 1-
 - 17
 - 18
 - 35
 - 35.5
- What element is represented by the Bohr-Rutherford diagram below?
 - potassium, K
 - calcium, Ca
 - hydrogen, H
 - argon, Ar
 - yttrium, Y
- Which of the following statements about the masses of atoms is *false*?
 - Most of the mass of an atom is located in the nucleus.
 - Electrons contribute very little to the mass of an atom.
 - Atoms of different elements have different masses.
 - Most of the mass of an atom's nucleus is due to the mass of neutrons.
 - The mass number of an atom is equal to the number of protons plus the number of neutrons.
- A neutral atom has eight electrons and eight neutrons. Which of the following options correctly represents the atom?
 - carbon-16
 - oxygen-16
 - oxygen-24
 - oxygen-8
 - sulfur-8
- Which of the following symbols represents the element with an atomic number of 14?
 - C
 - N
 - Ni
 - Si
 - S
- Which of the following statements is *false*?
 - The number of neutrons an atom has defines what element it is.
 - The number of electrons an atom has defines what element it is.
 - The number of protons an atom has defines what element it is.
 - All of the above are true.
 - None of the above are true.



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Table of Contents Mcgraw Hill Ryerson Science Answers

1. Understanding the eBook Mcgraw Hill Ryerson Science Answers
 - The Rise of Digital Reading Mcgraw Hill Ryerson Science Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Ryerson Science Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Ryerson Science Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Ryerson Science Answers
 - Personalized Recommendations
 - Mcgraw Hill Ryerson Science Answers User Reviews and Ratings
 - Mcgraw Hill Ryerson Science Answers and Bestseller Lists
5. Accessing Mcgraw Hill Ryerson Science Answers Free and Paid eBooks
 - Mcgraw Hill Ryerson Science Answers Public Domain eBooks
 - Mcgraw Hill Ryerson Science Answers eBook Subscription Services
 - Mcgraw Hill Ryerson Science Answers Budget-Friendly Options
6. Navigating Mcgraw Hill Ryerson Science Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcgraw Hill Ryerson Science Answers Compatibility with Devices
 - Mcgraw Hill Ryerson Science Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Ryerson Science Answers
 - Highlighting and Note-Taking Mcgraw Hill Ryerson Science Answers
 - Interactive Elements Mcgraw Hill Ryerson Science Answers
8. Staying Engaged with Mcgraw Hill Ryerson Science Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Ryerson Science Answers
9. Balancing eBooks and Physical Books Mcgraw Hill Ryerson Science Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Ryerson Science Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcgraw Hill Ryerson Science Answers
- Setting Reading Goals Mcgraw Hill Ryerson Science Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcgraw Hill Ryerson Science Answers
- Fact-Checking eBook Content of Mcgraw Hill Ryerson Science Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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