

## The formula of Epsom salt

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**Date performed:** September 10, 2014

**Date submitted:** September 17, 2014

**Aim:** Epsom salts are sold as health salts. They are hydrated magnesium sulphate. That means they contain water, called water of crystallization, as well as magnesium sulphate. The formula can be written as  $\text{MgSO}_4 \cdot x\text{H}_2\text{O}$ , and the aim is to find the value for  $x$  in the formula of Epsom salt.

When Epsom salts are heated, the water of crystallization is driven off leaving anhydrous magnesium sulphate,  $\text{MgSO}_4$ .

### **Chemicals and Equipment:**

Epsom salt, crucible, tripod, triangle, spatula, Bunsen burner, balance

### **Procedure**

Record all you weightings in a data table

1. Weight a clean and dry crucible
2. Add approximately 0.5 to 1 cm depth of Epsom salt and note the mass.
3. Place the crucible in the triangle on the tripod.
4. Heat gently at first then strongly for about 5 minutes.
5. Let the crucible cool to about 30 °C. Weigh.
6. Heat again for a couple of minutes. Allow to cool and weigh.
7. If the last two weightings differ by more than 0.05 g, repeat the heating and weighting.
8. Calculate a value for  $x$  in  $\text{MgSO}_4 \cdot x\text{H}_2\text{O}$ .

Mass of crucible: 30.24 g

Mass of crucible and salt: 37.03 g

Mass of salt: 6.79 g

	Before heating	After 1 <sup>st</sup> heating	After 2 <sup>nd</sup> heating	After 3 <sup>rd</sup> heating
Mass of substance (g ±0.01)	6.79	3.87	3.10	3.33
Mass difference in substance (g ±0.01)	0.00	-2.92	-0.77	+0.23

# Epsom Salt Crystal Lab Report

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