

Thermochemistry Lab Report: Calorimeter Constants, Hess's Law, and Enthalpy Changes

7 pages | 1666 words

Abstract

The aim of these three experiments was to determine the heat capacity of a calorimeter, confirm Hess's Law, and observe enthalpy changes within reactions. By measuring the change in temperature resulting from the interaction of various reactants, we determined both the calorimeter constant and the change in enthalpy for the reactions. The results varied, with some values closely resembling theoretical ones, while others showed greater discrepancies.

Introduction

The first experiment focuses on finding the calorimeter constant for a polystyrene cup. During chemical reactions inside a calorimeter, some heat is lost to the surroundings. To ensure accurate results in subsequent experiments, it is crucial to determine the amount of heat lost. The equation used to calculate this constant is derived from the overall heat of the reaction equation: Overall Heat = - [(Sp.Ht. hotwater * Mass of water * Change in temperature) + (Sp.Ht. coolwater * Mass of water * Change in temperature) + (Cp calorimeter * Change in temperature)] Since experimental errors are inevitable, three calculations were performed to obtain an average calorimeter constant. This experiment is vital for the success of the following two thermochemistry experiments. The second experiment, known as Hess's Law, aims to confirm this fundamental law. Three reactions are examined, where one of them is the same as the other two, and the heats of reaction for each are measured. Hess's Law states that the heat of reaction for one reaction should equal the sum of the heats of reaction for the other two reactions. The reactions used are:

1. $\text{NaOH}(s) \rightarrow \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$
2. $\text{NaOH}(s) + \text{HCl}(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$
3. $\text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$

To find the heat released by each reaction, a modified version of the overall heat of reaction equation is used: $q = - [\text{Sp.Ht.} * m * \Delta t]$ Additionally, changes in entropy (ΔS) are calculated using theoretical values from reference tables. Finally, the overall free energy change (ΔG) is determined using the equation: $\Delta G = \Delta H - (\text{Temperature} * \Delta S)$ All of this is used to verify Hess's Law by calculating the percent error involved in the experiment. The third experiment, titled "Thermochemistry: Acid + Base," combines the concepts of the previous two experiments. It observes the change in enthalpy resulting from various reactions between strong and weak acids and bases. Four reactions are investigated:

1. $\text{HCl}(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{NaCl}(\text{aq}) + \text{H}_2\text{O}(\text{l})$
2. $\text{HCl}(\text{aq}) + \text{NH}_3(\text{aq}) \rightarrow \text{NH}_4\text{Cl}(\text{aq})$
3. $\text{HC}_2\text{H}_3\text{O}_2(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{NaC}_2\text{H}_3\text{O}_2(\text{aq}) + \text{H}_2\text{O}(\text{l})$
4. $\text{HC}_2\text{H}_3\text{O}_2(\text{aq}) + \text{NH}_3(\text{aq}) \rightarrow \text{NH}_4\text{C}_2\text{H}_3\text{O}_2(\text{aq})$

By monitoring the change in temperature resulting from the reaction of an acid and a base, the overall energy change for each reaction ($\Delta H_{\text{rxn/mole}}$ of limiting reactant) is calculated. This experimental value is compared to the theoretical value to determine the accuracy of the experiment.

Experimental

In order to do any calculation for energy, we first had to find the calorimeter constant. In order to do that, we first took and weighed a polystyrene cup (our calorimeter) and added approximately 100 g of warm water to it. The actual measurements are recorded in...

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Table of Contents Enthalpy Calorimeter Lab Report

1. Understanding the eBook Enthalpy Calorimeter Lab Report
 - The Rise of Digital Reading Enthalpy Calorimeter Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Enthalpy Calorimeter Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Enthalpy Calorimeter Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Enthalpy Calorimeter Lab Report
 - Personalized Recommendations
 - Enthalpy Calorimeter Lab Report User Reviews and Ratings
 - Enthalpy Calorimeter Lab Report and Bestseller Lists

5. Accessing Enthalpy Calorimeter Lab Report Free and Paid eBooks
 - Enthalpy Calorimeter Lab Report Public Domain eBooks
 - Enthalpy Calorimeter Lab Report eBook Subscription Services
 - Enthalpy Calorimeter Lab Report Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Enthalpy Calorimeter Lab Report Compatibility with Devices
 - Enthalpy Calorimeter Lab Report Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Enthalpy Calorimeter Lab Report
 - Highlighting and Note-Taking Enthalpy Calorimeter Lab Report
 - Interactive Elements Enthalpy Calorimeter Lab Report
8. Staying Engaged with Enthalpy Calorimeter Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Enthalpy Calorimeter Lab Report
9. Balancing eBooks and Physical Books Enthalpy Calorimeter Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Enthalpy Calorimeter Lab Report
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Enthalpy Calorimeter Lab Report
 - Setting Reading Goals Enthalpy Calorimeter Lab Report
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Enthalpy Calorimeter Lab Report
 - Fact-Checking eBook Content of Enthalpy Calorimeter Lab Report
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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