

Gas Laws Practice - Combined, Individual, and Ideal Gas Laws (Answers on next page)

Ideal Gas Law:

$$PV = nRT$$

$$R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$$

$$R = 62.4 \frac{\text{L} \cdot \text{mmHg}}{\text{mol} \cdot \text{K}}$$

$$R = 8.314 \frac{\text{L} \cdot \text{kPa}}{\text{mol} \cdot \text{K}}$$

Combined Law:

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Important Conversions:

$$K = ^\circ C + 273$$

$$1 \text{ atm} = 760 \text{ mmHg} = 760 \text{ Torr} = 101.3 \text{ kPa}$$

Note: Torr is just another pressure unit that is equivalent to mmHg.

- It is not safe to put aerosol canisters in a campfire because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2.0 moles of gas, and the campfire temperature is 1400.^o C, what is the pressure in mmHg inside the canister?
- A sample of nitrogen gas is cooled from 1200.^o C. What is its final temperature (in K) if it compresses from 12.4 L to 2.50 liters?
- Zach has added some volume (in milliliters) of air to a balloon at sea level (1.0 atm). If I take the balloon with me to Denver where the air pressure is 0.85 atm, the volume of the balloon is 15.0 L. What was the original volume of the balloon?
- How many moles of gas are in a 30.0-liter scuba canister if the temperature of the canister is 300. K and the pressure is 20300 kPa?
- A small research submarine with a volume of 1.2×10^5 L has an internal pressure of 1.0 atm and an internal temperature of 15^o C. If the submarine descends to a depth where the pressure is 150 atm and the temperature is 3.0^o C, what would be the new volume (in L) of the gas if the hull of the submarine is crushed?
- If 35.0 grams of helium are held at 115 kPa and 120 liters, what is the temperature (in ^oC)?
- What pressure (in Torr) is required to compress 196.0 liters of ammonia (NH₃) at 1.00 atmosphere into a cylinder whose volume is 26.0 liters? Assume temperature stays constant.
- My car has a constant internal volume of 2,600 liters. If the sun heats my car from a temperature of 20.^oC to a temperature of 55 ^oC, what will the pressure inside my car be if the pressure was initially 760 mmHg?
- How many grams of gas are initially in my car if the gas is oxygen (O₂) in question #8?
- Consider a closed container of gas.
 - Why does the pressure of the gas increase if more gas is added to the container if all other variables are kept constant?
 - If the volume of the container is compressed, why does pressure of the gas increase if all other variables are kept constant?
 - If the volume of the container is compressed, why does temperature of the gas decrease if all other variables are kept constant?
 - If the gas is cooled, why does pressure of the gas decrease if all other variables are kept constant?

Gas Laws Practice Calculations Answers Unit Lesson 01

Michael Hoelscher



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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

Resources in Education, 1973

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Chemical Principles Monty J. Montjar, 1990

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