

NAME: _____

DATE: _____

PERIOD: _____

GAS STOICHIOMETRY PROBLEMS WORKSHEET 1

1. Given the unbalanced decomposition reaction of baking soda:



42.0 grams of baking soda? What volumes of carbon dioxide and water are produced at STP?

$$42.0 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol CO}_2}{2 \text{ mol NaHCO}_3} \times \frac{22.4 \text{ L CO}_2}{1 \text{ mol CO}_2} = \boxed{5.6 \text{ L CO}_2}$$

$$42.0 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol NaHCO}_3} \times \frac{22.4 \text{ L H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{5.6 \text{ L H}_2\text{O}}$$

2. The catalytic decomposition of hydrogen peroxide is:



Balance the reaction. How many moles of water and oxygen are produced by the decomposition of 68.0 grams of hydrogen peroxide? How many molecules of water and oxygen are produced? How many grams of each product are formed?

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} = \boxed{2.00 \text{ mol H}_2\text{O}}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} = \boxed{1.00 \text{ mol O}_2}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} \times \frac{6.022 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{1.2044 \times 10^{24} \text{ molecules H}_2\text{O}}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{6.022 \times 10^{23} \text{ molecules O}_2}{1 \text{ mol O}_2} = \boxed{6.022 \times 10^{23} \text{ molecules O}_2}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{36.04 \text{ g H}_2\text{O}}$$

$$68 \text{ g H}_2\text{O}_2 \times \frac{1 \text{ mol H}_2\text{O}_2}{34.01 \text{ g H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{32.00 \text{ g O}_2}{1 \text{ mol O}_2} = \boxed{63.96 \text{ g O}_2}$$

3. If 156.0 grams of potassium metal reacts with excess water, potassium hydroxide and hydrogen gas are formed? What volume of hydrogen gas, in liters,

Be sure to balance the reaction first.



$$156 \text{ g K} \times \frac{1 \text{ mol K}}{39.09 \text{ g K}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol K}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = \boxed{44.70 \text{ L H}_2}$$

$$156 \text{ g K} \times \frac{1 \text{ mol K}}{39.09 \text{ g K}} \times \frac{2 \text{ mol KOH}}{2 \text{ mol K}} \times \frac{56.11 \text{ g KOH}}{1 \text{ mol KOH}} = \boxed{87.39 \text{ g KOH}}$$

4. Determine the number of moles of carbon dioxide gas, water, and sodium chloride formed by the reaction of 42.0 grams of sodium bicarbonate (baking soda) reacting with excess hydrochloric acid. The reaction is:



$$42 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol NaHCO}_3} = \boxed{0.5 \text{ mol CO}_2}$$

$$42 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol NaHCO}_3} = \boxed{0.5 \text{ mol H}_2\text{O}}$$

$$42 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol NaCl}}{1 \text{ mol NaHCO}_3} = \boxed{0.5 \text{ mol NaCl}}$$

Gas Stoichiometry Worksheet Answer Key With Work

Mike Jess



Gas Stoichiometry Worksheet Answer Key With Work:

Reviewing **Gas Stoichiometry Worksheet Answer Key With Work**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Gas Stoichiometry Worksheet Answer Key With Work**," an enthralling opus penned by a highly acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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