

13. Consider the following statements
- Atmospheric pressure is less at the top of a mountain than at sea level
 - Gases are much more compressible than solids or liquids
 - When the atmospheric pressure increases the height of the mercury column rises
- Select the correct statement
- I and II
 - II and III
 - I and III
 - I, II and III
14. Compressibility factor for CO_2 at 400 K and 71.0 bar is 0.8697. The molar volume of CO_2 under these conditions is
- 22.04 dm^3
 - 2.24 dm^3
 - 0.41 dm^3
 - 19.5 dm^3
15. If temperature and volume of an ideal gas is increased to twice its values, the initial pressure P becomes
- 4P
 - 2P
 - P
 - 3P
16. At identical temperature and pressure, the rate of diffusion of hydrogen gas is $3\sqrt{3}$ times that of a hydrocarbon having molecular formula $\text{C}_n\text{H}_{2n-2}$. What is the value of n ?
- 8
 - 4
 - 3
 - 1
17. Equal moles of hydrogen and oxygen gases are placed in a container, with a pin-hole through which both can escape what fraction of oxygen escapes in the time required for one-half of the hydrogen to escape. (NEET phase I)
- $\frac{3}{8}$
 - $\frac{1}{2}$
 - $\frac{1}{8}$
 - $\frac{1}{4}$
18. The variation of volume V , with temperature T , keeping pressure constant is called the coefficient of thermal expansion i.e. $\alpha = \frac{1}{V} \left(\frac{\partial V}{\partial T} \right)_P$. For an ideal gas α is equal to
- T
 - $\frac{1}{T}$
 - P
 - none of these
19. Four gases P, Q, R and S have almost same values of 'b' but their 'a' values (a, b are Vander Waals Constants) are in the order $Q < R < S < P$. At a particular temperature, among the four gases the most easily liquefiable one is
- P
 - Q
 - R
 - S
20. Maximum deviation from ideal gas is expected from (NEET)
- $\text{CH}_4(\text{g})$
 - $\text{NH}_3(\text{g})$
 - $\text{H}_2(\text{g})$
 - $\text{N}_2(\text{g})$
21. The units of Vander Waals constants 'b' and 'a' respectively
- mol L^{-1} and $\text{L atm}^2 \text{mol}^{-2}$
 - mol L and L atm mol^2
 - $\text{mol}^{-1} \text{L}$ and $\text{L}^2 \text{atm mol}^{-2}$
 - none of these

Gaseous State Multiple Choice Questions

Joacim Rocklöv



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Table of Contents Gaseous State Multiple Choice Questions

1. Understanding the eBook Gaseous State Multiple Choice Questions
 - The Rise of Digital Reading Gaseous State Multiple Choice Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Gaseous State Multiple Choice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Gaseous State Multiple Choice Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gaseous State Multiple Choice Questions
 - Personalized Recommendations
 - Gaseous State Multiple Choice Questions User Reviews and Ratings

- Gaseous State Multiple Choice Questions and Bestseller Lists
- 5. Accessing Gaseous State Multiple Choice Questions Free and Paid eBooks
 - Gaseous State Multiple Choice Questions Public Domain eBooks
 - Gaseous State Multiple Choice Questions eBook Subscription Services
 - Gaseous State Multiple Choice Questions Budget-Friendly Options
- 6. Navigating Gaseous State Multiple Choice Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Gaseous State Multiple Choice Questions Compatibility with Devices
 - Gaseous State Multiple Choice Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gaseous State Multiple Choice Questions
 - Highlighting and Note-Taking Gaseous State Multiple Choice Questions
 - Interactive Elements Gaseous State Multiple Choice Questions
- 8. Staying Engaged with Gaseous State Multiple Choice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gaseous State Multiple Choice Questions
- 9. Balancing eBooks and Physical Books Gaseous State Multiple Choice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gaseous State Multiple Choice Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gaseous State Multiple Choice Questions
 - Setting Reading Goals Gaseous State Multiple Choice Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gaseous State Multiple Choice Questions
 - Fact-Checking eBook Content of Gaseous State Multiple Choice Questions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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