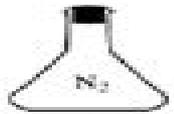
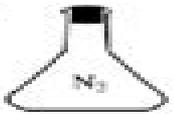
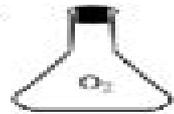
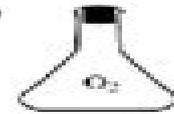
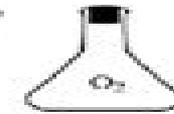


## Unit 2 Gases Study Guide

Name: \_\_\_\_\_

### Multiple Choice

Questions 1-3 refer to the 1 L flasks shown below. Provide an explanation for your answer.

(A)		(B)		(C)		(D)		(E)	
	T = 30°C P = 1 atm		T = 50°C P = 0.5 atm		T = 40°C P = 2 atm		T = 50°C P = 1 atm		T = 40°C P = 0.5 atm

- In which flask do the molecules have the greatest average speed?  
(A) A  
(B) B  
(C) C  
(D) D  
(E) E
- Which flask contains the smallest amount of moles?  
(A) A  
(B) B  
(C) C  
(D) D  
(E) E
- Which flask contains the sample with the greatest density?  
(A) A  
(B) B  
(C) C  
(D) D  
(E) E
- Under which of the following conditions of temperature and pressure will  $H_2$  gas be expected to behave most like an ideal gas?  
(A) 50 K and 0.10 atm  
(B) 50 K and 5.0 atm  
(C) 500 K and 0.10 atm

# Gases Multiple Choice Questions

**Marcel A. Müller**



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