

Unit 6 - Chemical Kinetics

Rate of reaction: the change in concentration of reactants or products per unit time.

Unit: $\text{mol dm}^{-3} \text{ s}^{-1}$

Methods for measuring rate of reaction	
Change in volume method	Change in mass method
If <u>one of the products is a gas</u> , a gas syringe can be connected to measure the increase in volume of gas per unit time. This can be achieved by measuring the volume every 30 seconds and plotting a graph of volume against time.	If there is a <u>change in mass of the reactants</u> , such as when a gas is evolved, the decrease in mass can be measured. This can be achieved by placing the reaction on a scale and weighing it per unit time.
Change in transmission of light method	Change in concentration method
If <u>one of the reactants or products is colored</u> , or an indicator is added to the reaction, a colorimeter or photo spectrophotometer can measure the change in intensity of light.	<u>Titration</u> can be used to measure the concentration of reactants or products per unit time, by titrating a known volume and concentration of a substance against the product or reactant. However, titration takes time, so a process known as <u>quenching</u> is used to stop the reaction and one can then measure the concentration of the products/reactants.
Change in conductivity method	Non-continuous methods
Using <u>electrolysis</u> , we measure the total electrical conductivity of a solution. As reactant ions are converted into product compounds, the decrease in total electrical conductivity will tell us the change in the concentration	By measuring the <u>time a reaction takes to reach a certain end point</u> under different reaction conditions, we are able to determine the rate of the reaction. However, this rate of reaction is only an average rate and not a continuous rate

Ib Chemistry Kinetics Study Guide

M Woodhall



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