

# Galvanic Cells

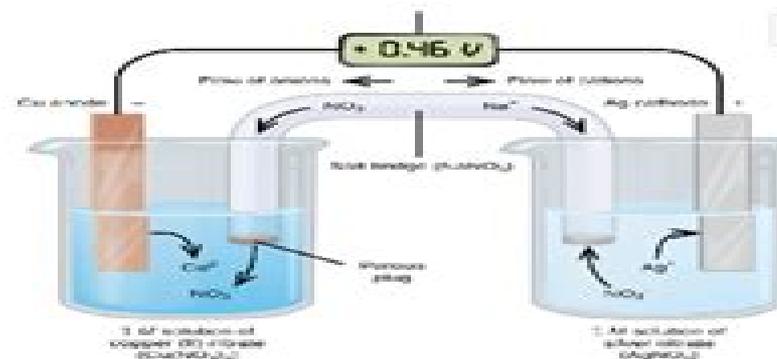
## 1. Introduction

**Aim** : The aim of this experiment is to measure the electric potential of a galvanic cell and compare it with the actual potential.

Galvanic cells are electrochemical cells in which electricity generates spontaneous oxidation reduction processes, commonly known as voltaic cells. When drafting the equations, oxidation - reduction reactions are typically separated into half-reactions to enable the entire equation to be balanced and to stress the actual chemical transitions. Oxidation is the loss of electrons or an increase in the oxidation number of an atom. Reduction is the gain of electrons or a decrease in the oxidation number of an atom.

The cell is made up of two half cells. The half cell shown at the left of the figure contains copper in an aqueous solution of copper nitrate. The half cell on the right contains silver ions in an aqueous solution of silver nitrate. A tube containing an inert electrolyte solution, known as a **salt bridge** is used to separate the two half cells. The electrodes in an electrochemical cell are called the cathode and anode. They are made of a solid metal and are submerged into a solution of the electrodes' ion called the electrolyte. For example, in Figure 1 since the electrode is made up of copper, the electrolyte is in a solution of an  $\text{Cu}^{2+}$  ions. **The anode** is where the oxidation occurs. It is always drawn on the left. **The cathode** is where the reduction takes place. It is always drawn on the right. An external conductor is connected to both the electrodes which completes the circuit.

FIGURE 1. Schematics of a galvanic cell.



As oxidation takes place in the anode, Zn oxidises and turns into  $\text{Zn}^{2+}$

# Galvanic Cell Lab Report

**SJ Ball**



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