

Gas Law Stoichiometry Worksheet

Directions: Use significant figures and units in the problems below.

$PV = nRT$

SHOW ALL WORK!!!

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:



- a. What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal.

$4.81 \text{ g Na} \left| \frac{1 \text{ mol}}{23.0 \text{ g Na}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol}} \right| = 2.34 \text{ L Cl}_2$

2.34 L Cl_2



- a. Balance the above reaction.
 b. What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

$n = x$
 $V = O_2(x)$

$T_1 = 25 + 273 = 298 \text{ K}$
 $P_1 = 1.05 \text{ atm}$

$5.53 \text{ g C}_3\text{H}_8 \left| \frac{1 \text{ mol}}{44 \text{ g}} \right| \left| \frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} \right| = 0.628 \text{ moles}$
 $V = 14.6 \text{ L}$

14.6 L

3. Potassium permanganate, KMnO_4 , is produced commercially by oxidizing aqueous potassium manganate, K_2MnO_4 .



- a. Balance the above reaction.
 b. What volume of $\text{Cl}_2(g)$, measured at STP, is needed to produce 10.0 g of KMnO_4 ?

$V = \frac{nRT}{P}$

$10.0 \text{ g KMnO}_4 \left| \frac{1 \text{ mol}}{158 \text{ g}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol Cl}_2} \right| = 0.709 \text{ L Cl}_2$

0.709 L

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated.



- a. Balance the above reaction.
 b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of ammonia gas would be collected at 24°C and 752 mmHg?

$10.3 \text{ g Mg}_3\text{N}_2 \left| \frac{1 \text{ mol}}{100.9 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{1 \text{ mol}} \right| = 0.204 \text{ mol}$
 Lim Reagent

$10.3 \text{ g H}_2\text{O} \left| \frac{1 \text{ mol}}{18 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{3 \text{ mol}} \right| = 0.381 \text{ mol}$

$24 + 273 = 297 \text{ K}$
 $752 / 760 = 0.989 \text{ atm}$

5.03 L

$V = \frac{nRT}{P}$

$V = \frac{(0.204 \text{ mol} + 0.082 \text{ mol}) (297 \text{ K})}{0.989 \text{ atm}}$

5.03 L

Gas Stoichiometry Answer Key

Julia Schneider



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Table of Contents Gas Stoichiometry Answer Key

1. Understanding the eBook Gas Stoichiometry Answer Key
 - The Rise of Digital Reading Gas Stoichiometry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Stoichiometry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Stoichiometry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Stoichiometry Answer Key
 - Personalized Recommendations
 - Gas Stoichiometry Answer Key User Reviews and Ratings
 - Gas Stoichiometry Answer Key and Bestseller Lists
5. Accessing Gas Stoichiometry Answer Key Free and Paid eBooks
 - Gas Stoichiometry Answer Key Public Domain eBooks
 - Gas Stoichiometry Answer Key eBook Subscription Services
 - Gas Stoichiometry Answer Key Budget-Friendly Options
6. Navigating Gas Stoichiometry Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Stoichiometry Answer Key Compatibility with Devices
 - Gas Stoichiometry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Stoichiometry Answer Key
 - Highlighting and Note-Taking Gas Stoichiometry Answer Key
 - Interactive Elements Gas Stoichiometry Answer Key
8. Staying Engaged with Gas Stoichiometry Answer Key

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Stoichiometry Answer Key
9. Balancing eBooks and Physical Books Gas Stoichiometry Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Stoichiometry Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Stoichiometry Answer Key
- Setting Reading Goals Gas Stoichiometry Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Stoichiometry Answer Key
- Fact-Checking eBook Content of Gas Stoichiometry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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