

GRAM FORMULA MASS

Name Lia Kerner

Determine the gram formula mass (the mass of one mole) of each compound below.

$$1. \text{KMnO}_4 \quad \underline{158 \text{ g}}$$

$$39 + 55 + 4(16) = 158$$

$$2. \text{KCl} \quad \underline{74 \text{ g}}$$

$$39 + 35 = 74$$

$$3. \text{Na}_2\text{SO}_4 \quad \underline{142 \text{ g}}$$

$$2(23) + 32 + 4(16) = 142$$

$$4. \text{Ca(NO}_3)_2 \quad \underline{164 \text{ g}}$$

$$40 + 2(14 + 3 \times 16) = 164$$

$$40 + 2(62) = 164$$

$$5. \text{Al}_2(\text{SO}_4)_3 \quad \underline{342 \text{ g}}$$

$$2(27) + 3(32 + 4 \times 16) = 342$$

$$6. (\text{NH}_4)_3\text{PO}_4 \quad \underline{149 \text{ g}}$$

$$3(14 + 4 \times 1) + 31 + 4(16) = 149$$

$$7. \text{CuSO}_4 \cdot 5\text{H}_2\text{O} \quad \underline{250 \text{ g}}$$

$$64 + 96 + 5(2 + 16) = 250$$

$$8. \text{Mg}_3(\text{PO}_4)_2 \quad \underline{264 \text{ g}}$$

$$3(24) + (96) \times 2 = 264$$

$$9. 7n(\text{C}_6\text{H}_6\text{O}) \cdot 2\text{H}_2\text{O} \quad \underline{219 \text{ g}}$$

Gram Formula Mass Answer Key

James C. Hill, Bruce Edward Bursten

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