



Unit: Chemical Bonding **Answer Key** "Covalent Bonding" - WS #3

Directions: Please answer each question using your notes from today's class. ☺

1. What is the difference between a molecular formula, structural formula and an electron dot formula? Give an example of each.

A molecular formula gives the chemical make-up of one molecule. It does not show atom-to-atom connectivity. Example: C_2H_2

A structural formula uses lines instead of dots to show shared pairs of electrons.

Example: $H-H$

An electron dot formula uses dots instead of lines to show shared pairs of electrons.

Example: $H:H$

2. What is the difference between single, double, and triple covalent bonds? How many shared electrons are in each?

The difference between the types of bonds is: a single bond is two shared electrons, a double is four shared electrons, and a triple is six shared electrons.

3. Draw the electron dot and structural formulas for the following molecules.

Substance	Rough copy (work!)	Electron dot	Structural
a) Cl_2		<i>Note: The electron dot structures have dots in between each atom instead of lines.</i>	$\begin{array}{c} \cdot\cdot \\ \cdot \\ \text{Cl} \\ \cdot \\ \cdot\cdot \end{array} - \begin{array}{c} \cdot\cdot \\ \cdot \\ \text{Cl} \\ \cdot \\ \cdot\cdot \end{array}$
b) H_2S			$\begin{array}{c} \cdot\cdot \\ \cdot \\ \text{S} \\ \cdot \\ \cdot\cdot \\ \\ \text{H} \\ \cdot \\ \cdot\cdot \end{array} - \text{H}$
c) BrI			$\begin{array}{c} \cdot\cdot \\ \cdot \\ \text{I} \\ \cdot \\ \cdot\cdot \end{array} - \begin{array}{c} \cdot\cdot \\ \cdot \\ \text{Br} \\ \cdot \\ \cdot\cdot \end{array}$

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