

General Chemistry II Jasperse
Kinetics. Extra Practice Problems

General Types/Groups of problems:

Rates of Change in Chemical Reactions	p1	First Order Rate Law Calculations	P9
The look of concentration/time graphs	p2	Reaction Energy Diagrams, Activation Energy, Transition States...	P10
Rates: Average Rates, Determination of Rates from Stoichiometry and Changes of Other Chemicals	p3	Reaction Mechanisms, Intermediates...	P12
Reactant Order and Overall Reaction Order	P4	Mechanism Steps and Rate Laws	P13
Given a Rate Law, How much will rate change with change in concentration	P5	Catalysts	P14
Determining Reactant Orders from Actual Data	P6	Answers	P15
Actual Rate Law from Concentration/Rate Data	P7		
Determining Rate Constant from Actual Data. Where you yourself need to first deduce the rate law, then plug in the values to solve for "k".	P8		

Rates of Change in Chemical Reactions

- For the reaction $A + 3B \rightarrow 2C$, how does the rate of disappearance of B compare to the rate of production of C?
 - the rate of disappearance of B is 1/2 the rate of appearance of C
 - the rate of disappearance of B is 3/2 the rate of appearance of C
 - the rate of disappearance of B is 2/3 the rate of appearance of C
 - the rate of disappearance of B is 1/3 the rate of appearance of C
- For the reaction $2A + 3B \rightarrow 4C + 5D$, the rate of the reaction in terms of ΔA would be written as:
 - $-\Delta A/\Delta t$
 - $-1/2 \Delta A/\Delta t$
 - $+\Delta A/\Delta t$
 - $+1/2 \Delta A/\Delta t$
 - $-2 \Delta A/\Delta t$
- For the reaction $2A + 3B \rightarrow 4C + 5D$, the rate of the reaction in terms of ΔB would be written as:
 - $-\Delta B/\Delta t$
 - $+\Delta B/\Delta t$
 - $-1/3 \Delta B/\Delta t$
 - $+1/3 \Delta B/\Delta t$
 - $-3 \Delta B/\Delta t$
- For the reaction $2A + 3B \rightarrow 4C + 5D$, the rate of the reaction in terms of ΔC would be written as:
 - $+\Delta C/\Delta t$
 - $+4 \Delta C/\Delta t$
 - $+1/4 \Delta C/\Delta t$
 - $-4 \Delta C/\Delta t$
 - $-1/4 \Delta C/\Delta t$
- In the combustion of methane, $CH_4(g) + 2 O_2(g) \rightarrow CO_2(g) + 2 H_2O(g)$, which reactant has the greatest rate of disappearance?
 - CH_4
 - O_2
 - CO_2
 - H_2O
 - CH_4 and O_2 have the same rate of disappearance.

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