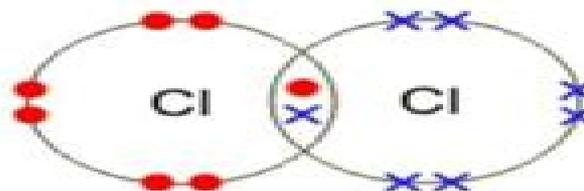


### Covalent bonding

A non-metal metal and another non-metal can bond together by **sharing** electrons so that both electrons can achieve a **full outer shell**. This is called **covalent bonding**. To represent the atoms involved in the bonding we draw crosses for the electrons on one of the atoms and dots for the electrons on the other atoms as shown here.



### Instructions

NB: It may be easiest to draw each atom out individually first to work out how many electrons will be shared

1. Draw the circles (outer shells) and overlap them (like a Venn diagram)
2. Write the symbols of the atoms involved in the middle of each circle
3. Draw the electrons for one atom as dots (o) and electrons for the other atom as crosses (x).
4. Any shared electrons are drawn in the **overlapping** section.
5. Count up the electrons for EACH atom and check that they each now have a **full outer shell** (2 electrons for the first shell and 8 for the next 2 shells)

**TASK:** Draw dot and cross diagrams for the following covalent compounds.

$F_2$		$CH_4$	
$HCl$		$CO_2$	
$H_2S$		$H_2O$	
$HF$		$NH_3$	

**Challenge** -  $O_2$

# Ionic And Metallic Bonding Questions

**Mina Lebitz**



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