

Ionic Bonds Practice

1. Fill in the missing information on the chart.

Na
Cl
Be
F
Li
O
P

Element	# of Protons	# of Electrons	# of Valence Electrons
Sodium	11	11	1
Chlorine	17	17	7
Beryllium	4	4	2
Fluorine	9	9	7
Lithium	3	3	1
Oxygen	8	8	6
Phosphorus	15	15	5

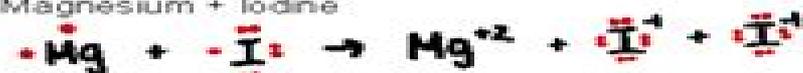
2. For each of the following ionic bonds:

- Write the symbols for each element.
- Draw a Lewis Dot structure for the valence shell of each element.
- Draw an arrow (or more if needed) to show the transfer of electrons to the new element.
- Write the charges on the ions.
- Write the resulting chemical formula.

a) Sodium + Chlorine



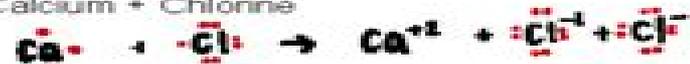
b) Magnesium + Iodine



c) Sodium + Oxygen



d) Calcium + Chlorine



e) Aluminum + Chlorine



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