

Physics

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(Chapter 2) (Units and Measurement) (Exercises)

(Class 11)

Question 2.15:

A famous relation in physics relates 'moving mass' m to the 'rest mass' m_0 of a particle in terms of its speed v and the speed of light, c . (This relation first arose as a consequence of special relativity due to Albert Einstein). A boy recalls the relation almost correctly but forgets where to put the constant c . He writes:

$$m = \frac{m_0}{(1 - v^2)^{1/2}}$$

Guess where to put the missing c .

Answer 2.15:

Given the relation,

$$m = \frac{m_0}{(1 - v^2)^{1/2}}$$

Dimension of $m = M^1 L^0 T^0$,

Dimension of $v = M^0 L^1 T^{-1}$,

Dimension of $c = M^0 L^1 T^{-1}$

Dimension of $m_0 = M^1 L^0 T^0$

Dimension of $v^2 = M^0 L^2 T^{-2}$

The given formula will be dimensionally correct only when the dimension of L.H.S is the same as that of R.H.S.

This is only possible when the factor, $(1 - v^2)^{1/2}$ is dimensionless i.e., $(1 - v^2)$ is dimensionless.

This is only possible if v^2 is divided by c^2 .

Hence, the correct relation is

$$m = \frac{m_0}{\left(1 - \frac{v^2}{c^2}\right)^{1/2}}$$

Question 2.16:

The unit of length convenient on the atomic scale is known as an angstrom and is denoted by \AA : $1 \text{\AA} = 10^{-10} \text{ m}$. The size of a hydrogen atom is about 0.5\AA . What is the total atomic volume in m^3 of a mole of hydrogen atoms?

Answer 2.16:

Radius of hydrogen atom, $r = 0.5 \text{\AA} = 0.5 \times 10^{-10} \text{ m}$

Volume of hydrogen atom = $\frac{4}{3} \pi r^3$

$$= \frac{4}{3} \times \frac{22}{7} \times (0.5 \times 10^{-10})^3 = 0.524 \times 10^{-30} \text{ m}^3$$

1 mole of hydrogen contains 6.023×10^{23} hydrogen atoms.

Volume of 1 mole of hydrogen atoms = $6.023 \times 10^{23} \times 0.524 \times 10^{-30}$

$$= 3.16 \times 10^{-7} \text{ m}^3$$

Question 2.17:

One mole of an ideal gas at standard temperature and pressure occupies 22.4 L (molar volume). What is the ratio of molar volume to the atomic volume of a mole of hydrogen? (Take the size of hydrogen molecule to be about 1\AA). Why is this ratio so large?

Answer 2.17:

Radius of hydrogen atom, $r = 0.5 \text{\AA} = 0.5 \times 10^{-10} \text{ m}$

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Table of Contents Grade 11 Physcs Paper 2

1. Understanding the eBook Grade 11 Physcs Paper 2
 - The Rise of Digital Reading Grade 11 Physcs Paper 2
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 11 Physcs Paper 2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 11 Physcs Paper 2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 11 Physcs Paper 2
 - Personalized Recommendations
 - Grade 11 Physcs Paper 2 User Reviews and Ratings
 - Grade 11 Physcs Paper 2 and Bestseller Lists

5. Accessing Grade 11 Physcs Paper 2 Free and Paid eBooks
 - Grade 11 Physcs Paper 2 Public Domain eBooks
 - Grade 11 Physcs Paper 2 eBook Subscription Services
 - Grade 11 Physcs Paper 2 Budget-Friendly Options
6. Navigating Grade 11 Physcs Paper 2 eBook Formats
 - ePub, PDF, MOBI, and More
 - Grade 11 Physcs Paper 2 Compatibility with Devices
 - Grade 11 Physcs Paper 2 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 11 Physcs Paper 2
 - Highlighting and Note-Taking Grade 11 Physcs Paper 2
 - Interactive Elements Grade 11 Physcs Paper 2
8. Staying Engaged with Grade 11 Physcs Paper 2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 11 Physcs Paper 2
9. Balancing eBooks and Physical Books Grade 11 Physcs Paper 2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 11 Physcs Paper 2
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Grade 11 Physcs Paper 2
 - Setting Reading Goals Grade 11 Physcs Paper 2
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade 11 Physcs Paper 2
 - Fact-Checking eBook Content of Grade 11 Physcs Paper 2
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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