

PRACTICE QUESTIONS TO INTRODUCTION TO ORGANIC CHEMISTRY
DATE : OCTOBER, 2023



INDATWA N' INKESHA SCHOOL

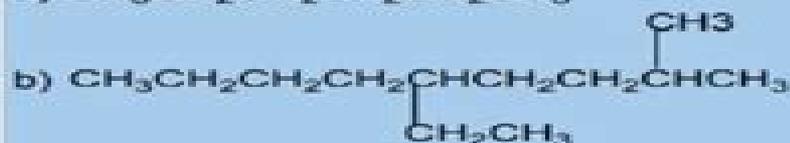
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G.S.O.B

P

Part I: Classification of hydrocarbon and determination of E.F and M.F

Q1. Observe the following compounds and classify them as aliphatic, alicyclic and aromatic.



Q2. Determine the empirical formula of a compound with the following composition by mass: 48.0 % C, 8.0 % H, 28.0 % N and 16.0 % O. If this compound has a molar mass of 200 g, what is its molecular formula?

Q3. Determine the empirical formula of a compound with the following composition by mass: 60.0 % C, 8.0 % H and 32.0 % O. If 100 mg of this compound is neutralized with exactly 10.0 mL of 0.0500 M NaOH in a reaction with a 1:1 stoichiometry, what is the molecular formula of this compound?

Q4. Combustion of 1.000 grams of an organic compound known to contain only C, H and O produces 2.360 g CO_2 and 0.640 g H_2O . What is the empirical formula of the compound?

Q5. Complete combustion of 1.110 g of a gaseous hydrocarbon yields 3.613g CO_2 , and 1.109g H_2O . A 0.288-g sample of the hydrocarbon occupies a volume of 131 mL at 24.8°C and 753mmHg. What is the molecular formula of the hydrocarbon? $R = 0.08206\text{Latm/mol.k}$, $1\text{atm} = 760\text{mmHg}$.

Q6. 20.0cm^3 of a gaseous, unsaturated hydrocarbon A were exploded with 200cm^3 of oxygen and, after cooling to the original room temperature, 140cm^3 of gas remained. This volume was 20.0cm^3 on shaking the gas with potassium hydroxide solution. Deduce the molecular formula of A.

Hydrocarbon Chemistry Exercises

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