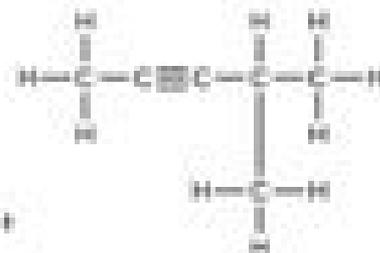
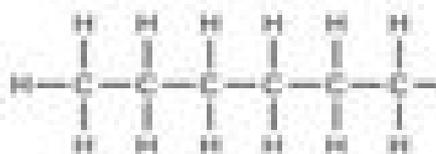


TOPIC 1 -

Hydrocarbons



Organic compounds containing only hydrogen and carbon atoms.

Key Facts

- Hydrocarbons may have one or more branched chains or cyclic structures.
- The hydrogen and carbon atoms have similar electronegativities (2.20 and 2.55) making them have low solubility in polar solvents.
- Hydrocarbons have the weakest intermolecular forces (London dispersion forces).
- Weak intermolecular forces → low boiling and melting points.
- 3 main groups of hydrocarbon
 - Alkanes ⇒ single carbon-carbon bonds (saturated hydrocarbons)
 - Alkenes ⇒ contain at least one carbon-carbon double bond (unsaturated hydrocarbons)
 - Alkynes ⇒ contain at least one carbon-carbon triple bond (unsaturated)

3 Main Groups of Hydrocarbon

Alkanes

- General formula = $\text{C}_n\text{H}_{2n+2}$
- Hydrocarbons with single

Alkenes

- General formula = C_nH_{2n}
- Hydrocarbons with at least

Alkynes

- General formula = $\text{C}_n\text{H}_{2n-2}$
- Hydrocarbons with at least

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